

All answers in cells and lines are valued at 1 pt.  
each unless marked otherwise.

Name \_\_\_\_\_ Period \_\_\_\_\_

# Atoms and Bonds Chapter Test

## Section 1 Fill in the Table

Compound/Molecule	Metal + Nonmetal or Nonmetal + Nonmetal?	Ionic or Covalent Bond?	e <sup>-</sup> shared or transferred?
KI			
Br <sub>2</sub>			
CO <sub>2</sub>			
PbF <sub>2</sub>			

## Section 2 Periodic Symbol Analysis

28
Ni
58.69

Element name \_\_\_\_\_  
 Name English or Latin in origin? \_\_\_\_\_  
 Atomic number \_\_\_\_\_ Number of protons \_\_\_\_\_  
 Atomic mass \_\_\_\_\_ Number of electrons \_\_\_\_\_  
 Number of neutrons (Mass - # protons) \_\_\_\_\_  
 Metal, nonmetal, or noble gas? \_\_\_\_\_

## Section 3 Properties of Bonds Write "I" if the property describes an ionic bond, "C" if the property describes a covalent bond, and "B" if the property describes both kinds of bonds.

\_\_\_\_\_ High melting points \_\_\_\_\_ Weaker bond \_\_\_\_\_ Electrons shared  
 \_\_\_\_\_ Conducts electricity in water \_\_\_\_\_ Helps most atoms achieve the octet of a noble gas  
 \_\_\_\_\_ Polar or nonpolar molecules \_\_\_\_\_ Attraction of opposite charges  
 \_\_\_\_\_ Poor conductors of electricity \_\_\_\_\_ Crystal shape \_\_\_\_\_ Low melting points  
 \_\_\_\_\_ Attraction between a cation and an anion

## Section 4 Formula Writing Write the formula for each of the compounds below. Be sure to balance the charges with subscripts. (Hint: Parenthesized Roman numerals are the charges of transition metals)

- Silver sulfide \_\_\_\_\_
- Nickel (II) cyanide \_\_\_\_\_
- Aluminum phosphate \_\_\_\_\_
- Aluminum bromide \_\_\_\_\_
- Copper (I) sulfate \_\_\_\_\_
- Silver bromide \_\_\_\_\_

Cyanide = CN<sup>-</sup> Sulfide = S<sup>2-</sup> Phosphate = PO<sub>4</sub><sup>3-</sup> Silver = Ag<sup>+</sup> Sulfate = SO<sub>4</sub><sup>2-</sup> Aluminum = Al<sup>3+</sup>

# Ionic Bonds Chapter Test Answer Key

**Tao Wei**



## **Ionic Bonds Chapter Test Answer Key:**

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## **Table of Contents Ionic Bonds Chapter Test Answer Key**

1. Understanding the eBook Ionic Bonds Chapter Test Answer Key
  - The Rise of Digital Reading Ionic Bonds Chapter Test Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds Chapter Test Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic Bonds Chapter Test Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds Chapter Test Answer Key
  - Personalized Recommendations
  - Ionic Bonds Chapter Test Answer Key User Reviews and Ratings
  - Ionic Bonds Chapter Test Answer Key and Bestseller Lists
5. Accessing Ionic Bonds Chapter Test Answer Key Free and Paid eBooks
  - Ionic Bonds Chapter Test Answer Key Public Domain eBooks
  - Ionic Bonds Chapter Test Answer Key eBook Subscription Services
  - Ionic Bonds Chapter Test Answer Key Budget-Friendly Options
6. Navigating Ionic Bonds Chapter Test Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Ionic Bonds Chapter Test Answer Key Compatibility with Devices
  - Ionic Bonds Chapter Test Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ionic Bonds Chapter Test Answer Key
  - Highlighting and Note-Taking Ionic Bonds Chapter Test Answer Key
  - Interactive Elements Ionic Bonds Chapter Test Answer Key

8. Staying Engaged with Ionic Bonds Chapter Test Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ionic Bonds Chapter Test Answer Key
9. Balancing eBooks and Physical Books Ionic Bonds Chapter Test Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ionic Bonds Chapter Test Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ionic Bonds Chapter Test Answer Key
  - Setting Reading Goals Ionic Bonds Chapter Test Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic Bonds Chapter Test Answer Key
  - Fact-Checking eBook Content of Ionic Bonds Chapter Test Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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