

### Question 1

(a) Using only the periodic table, arrange each set of atoms in order of increasing radius:

(1) Ba, Ca, Na



(2) In, Sn, As



(3) Cs, K, Al



(b) Which neutral atom is isoelectronic with each of the following ions?

(1)  $\text{H}^+$

None

(The  $\text{H}^+$  ion has 0 electrons. It is impossible for a neutral atom to have 0 electrons.)

(2)  $\text{Zr}^{4+}$

Kr

(3)  $\text{Mn}^{7+}$

Ar

(4)  $\text{I}^-$

Xe

(c) Explain the following variations in atomic or ionic radii:

(1)  $\text{I}^- > \text{I} > \text{I}^+$

As the atomic number stays constant and the number of electrons increases, the electron-electron repulsions increase. This leads to the electrons spreading out, and the anion becomes larger. The reverse is also true, so the cation is smaller.

(2)  $\text{Ca}^{2+} > \text{Mg}^{2+} > \text{Be}^{2+}$

For cations with the same charge, ionic radii increase going down a column because of the increase in principal quantum number, which means the outer electrons are farther away from the nucleus on average. This leads to larger ion size.

# Ionic Radius Answer Key

**Arshad Iqbal**



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