



# CHEMICAL BONDS



## IONIC, COVALENT, & METALLIC

**FORMS OF CHEMICAL BONDS**  
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

**Types of Bonds:**

- Covalent Bond:** Formed by sharing electrons between two atoms. Example:  $H_2$ ,  $H_2O$ ,  $CO_2$ .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, resulting in oppositely charged ions that attract each other. Example:  $NaCl$ ,  $MgO$ .
- Metallic Bond:** Formed by the attraction between metal cations and a sea of delocalized electrons. Example:  $Fe$ ,  $Cu$ ,  $Al$ .

**Covalent, Ionic, and Metallic Bonding**

Covalent bonding occurs between nonmetals. In this type of bonding, atoms share their outermost electrons to achieve a stable electron configuration. This sharing is represented by pairs of dots (electron pairs) between the atoms.

Ionic bonding occurs between a metal and a nonmetal. The metal atom loses one or more electrons to become a positively charged cation, and the nonmetal atom gains those electrons to become a negatively charged anion. The resulting ions are attracted to each other by electrostatic forces.

Metallic bonding occurs in metals. Metal atoms release their outermost electrons into a common pool, called the "sea of electrons." The metal cations are held together by the attraction of this electron sea.

**Science Real**

**Diagram of Ionic Bonding:** Shows a sodium atom ( $Na$ ) losing an electron to become a sodium cation ( $Na^+$ ) and a chlorine atom ( $Cl$ ) gaining an electron to become a chloride anion ( $Cl^-$ ). The resulting ions are attracted to each other by electrostatic forces, forming an ionic bond.

**Diagram of Covalent Bonding:** Shows two hydrogen atoms ( $H$ ) sharing their outermost electrons to form a hydrogen molecule ( $H_2$ ). The shared electrons are represented by a pair of dots between the two atoms.

**Diagram of Metallic Bonding:** Shows a lattice of metal cations held together by a sea of delocalized electrons.

**Diagram of Bond Types:** Shows the relative strength of different types of chemical bonds. The diagram indicates that ionic bonds are the strongest, followed by covalent bonds, and then metallic bonds.

**Diagram of Bond Formation:** Shows the process of bond formation between two atoms. It illustrates how the distance between the atoms changes as they approach each other, and how the potential energy of the system changes as the bond forms.

**Diagram of Bond Energy:** Shows the energy required to break a chemical bond. It indicates that the energy required to break a covalent bond is generally higher than the energy required to break an ionic bond.

**Diagram of Bond Length:** Shows the distance between the nuclei of two atoms in a molecule. It indicates that the bond length is generally shorter for covalent bonds than for ionic bonds.

**Diagram of Bond Angle:** Shows the angle between two bonds in a molecule. It indicates that the bond angle is generally smaller for covalent bonds than for ionic bonds.

**Diagram of Bond Polarity:** Shows the distribution of electron density in a bond. It indicates that covalent bonds can be polar or nonpolar, depending on the electronegativity of the atoms involved.

**Diagram of Bond Order:** Shows the number of electron pairs shared between two atoms. It indicates that single bonds are the most common, followed by double bonds, and then triple bonds.

**Diagram of Bond Dissociation Energy:** Shows the energy required to break a bond in a molecule. It indicates that the bond dissociation energy is generally higher for covalent bonds than for ionic bonds.

**Diagram of Bond Enthalpy:** Shows the heat change associated with the formation or breaking of a bond. It indicates that the bond enthalpy is generally higher for covalent bonds than for ionic bonds.

**Diagram of Bond Length and Bond Angle:** Shows the relationship between bond length and bond angle. It indicates that bond length and bond angle are related, and can be used to predict the shape of a molecule.

**Diagram of Bond Polarity and Bond Order:** Shows the relationship between bond polarity and bond order. It indicates that bond polarity and bond order are related, and can be used to predict the properties of a molecule.

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## GUIDED READING ACTIVITY

# Ionic And Metallic Bonding Guided Reading

**AN Whitehead**



## **Ionic And Metallic Bonding Guided Reading:**

## Embracing the Song of Phrase: An Emotional Symphony within **Ionic And Metallic Bonding Guided Reading**

In a global eaten by monitors and the ceaseless chatter of quick interaction, the melodic splendor and mental symphony produced by the prepared word usually fade in to the background, eclipsed by the relentless sound and disruptions that permeate our lives. But, located within the pages of **Ionic And Metallic Bonding Guided Reading** a charming fictional prize overflowing with natural emotions, lies an immersive symphony waiting to be embraced. Constructed by a wonderful composer of language, this captivating masterpiece conducts viewers on a psychological journey, skillfully unraveling the hidden songs and profound affect resonating within each carefully constructed phrase. Within the depths of the emotional assessment, we shall examine the book is key harmonies, analyze its enthralling publishing model, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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