

\_\_\_\_\_

Identify the order of the change that best completes the statement or answers the question.

6. In many compounds, atoms of main-group elements form ions so that the number of electrons in their outermost energy levels of each ion is:

  - a. 2.
  - b. 8.
  - c. 10.
  - d. 18.
7. An atom

  - a. is an ion with a negative charge.
  - b. attracts ions with negative charges.
  - c. results when an alkaline-earth metal loses one of its two outermost electrons.
  - d. has more protons than electrons.
8. When the octet rule is satisfied, the outermost \_\_\_\_\_ are filled.

  - a. all orbitals
  - b. s and p orbitals
  - c. d orbitals
  - d. s and p orbitals
9. The elements of the \_\_\_\_\_ group satisfy the octet rule without forming compounds.

  - a. halide
  - b. noble gas
  - c. alkali metal
  - d. alkaline-earth metal
10. Chlorine can have half p and p orbitals in its outermost energy level.

  - a. It is highly reactive only with alkali metals.
  - b. It is highly reactive only with halogens.
  - c. It can be combined with most elements.
  - d. It has a stable outer shell and is unreactive.
11. In a crystal of an ionic compound, each cation is surrounded by

  - a. molecules.
  - b. positive ions.
  - c. dipole.
  - d. anions.
12. The lithium(I) ion and sodium(I) ion

  - a. have the same charge.
  - b. are polycations ions.

# Ionic And Metallic Bonding Chapter 7 Worksheet Answers

**H Kauffman**



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