

Ch. 4 Guided Reading

4.1 Introduction to Atoms

Guided Reading and Study

Use Target Reading Skills

Sample questions and answers:

Q. What particles are in the center of an atom?

A. Protons and neutrons.

Q. What particles move around the outside of the nucleus?

A. Electrons.

1. False

2. a, c, d

3. True

4. J. J. Thomson

5. gold-foil experiment

6. Protons

7. Niels Bohr

8. Neutron

9. False

10. a, b

11. a. nucleus

b. neutron

c. proton

d. electron cloud

12. The number of protons in an atom equals

the number of electrons. The positive charge from the protons balances the negative charge from the electrons, making the atom neutral.

13. protons and neutrons

14. The mass of an electron is about 1/2,000 (Or 1/1,836) of the mass of a proton.

15. protons

16. The atomic number is the number of protons in the nucleus.

17. Isotopes are atoms with the same number of protons but with a different number of neutrons.

18. Student diagrams should indicate any two of the following: An isotope of carbon with a mass number of 12 has 6 protons, 6 neutrons, and 6 electrons. An isotope of carbon with a mass number of 13 has 6 protons, 7 neutrons, and 6 electrons. An isotope of carbon with a mass number of 14 has 6 protons, 8 neutrons, and 6 electrons.

4.2 Organizing the Elements

Guided Reading and Study

1. He recognized a set of patterns that applied to all elements.

Introduction To Atoms Guided Reading And Study Answers

Ying-Ying Zheng



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