Activity B:	Get the Gizmo ready:	0,00
Methane reaction	Select C + H ₂ becomes CH ₄ .	36.00

Goal: Predict the results of a chemical reaction.

- Count: CH₄ is the chemical formula for methane. In order to produce a single molecule of methane, how many carbon (C) atoms are needed? Hydrogen atoms?
- Predict: Set the number of C atoms to 14 and the number of H₂ molecules to 10.
 - A. How many total hydrogen atoms are there?

B. How many methane molecules could form from these reactants?

55

20

C. After the atoms react, which reactant will be left over? The carbon

D. Which reactant will be the limiting reactant?

Hydrogen

E. Click Play and wait until Reaction complete is shown. What happened?

Each hydrogen molecule formed a connection with carbon (4 hydrogen molecules percarbon). There are extra carbon atoms that did not link because there are additional carbon atoms.

Test: Click Reset. Set the number of C atoms to seven.

A. How many hydrogen molecules (H₂) will be needed for there to be no excess

reactants? 14 Use the slider to set the number of H, molecules to this value.

B. How many methane molecules will be formed?

150

4

C. Click Play. What happened?

Since all of the hydrogen molecules linked to the carbon. there were no extra reactants.

 Solve: For each chemical equation below, write the number of product molecules that will form from the reaction. Then, circle the limiting reactant. (Note: The coefficients in front of the reactants indicate the number of reactant molecules or atoms present.)

<u>Full Version Student Exploration Limiting Reactants</u> <u>Answer Key Gizmo</u>

S Ben Porath

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