

Note that the end vowel of the Greek prefix is occasionally omitted from the spelling for clarity in pronunciation.

Hydrates

Prefixes are also used for the naming of **hydrates**, inorganic salts in which water molecules are a part of the crystalline structure of the solid. The waters of hydration (also called waters of crystallization) are bound to the ions of the salt and, in most cases, can be removed with the application of heat (see *Experiment 5*). For example, barium chloride is often purchased as barium chloride dihydrate: $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$. The formula of iron(III) chloride hexahydrate is $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$. The formula for Epsom salt (Figure 5.1) is $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$, magnesium sulfate heptahydrate.

Binary Acids

A binary acid is an *aqueous* solution of a compound formed by hydrogen and a more electronegative nonmetal. To name the acid, the prefix *hydro-* and suffix *-ic* are added to the root of the nonmetal name.

| Formula | Name |
|------------------------------------|--------------------------|
| HCl(aq) | <i>hydrochloric acid</i> |
| HBr(aq) | <i>hydrobromic acid</i> |
| $\text{H}_2\text{SO}_3\text{(aq)}$ | <i>sulfurous acid</i> |

Writing Formulas from Names

In writing formulas for compounds, it is mandatory that the sum of the oxidation numbers of the elements equals zero (see *Dry Lab 2A, rule 8*). The name of the compound must dictate how the formula is to be written.

Example D2B.1 Write the formula for calcium nitride.

Calcium is in group 2A, and therefore has an oxidation number of $+2$: Ca^{2+} . The nitrogen atom is always -3 when combined with a metal: N^{3-} . Therefore, for the sum of the oxidation numbers to equal zero, there must be three Ca^{2+} (a total of $+6$) for every two N^{3-} (a total of -6); the formula is Ca_3N_2 .

DRY LAB PROCEDURE

Procedure Overview: Given the formula of the compound, the proper names for a large number of compounds are to be written. Given the name of the compound, the formulas for a large number of compounds also are to be written.

Your instructor will assign the exercises you are to complete. Answer them on a separate piece of paper. Be sure to indicate the date, your lab section, and your desk number on your *Report Sheet*. Use the rules that have been described.

1. Name the following binary salts of the representative elements.

- a. Na_3P d. CaC_2 g. Ca_3P_2 j. K_2S m. NH_4Br p. AlCl_3
 b. Na_2O e. CaI_2 h. KCN k. K_2Te n. $(\text{NH}_4)_2\text{S}$ q. Al_2O_3
 c. Na_3N f. CaH_2 i. KOH l. K_2O_2 o. NH_4CN r. AlN

2. Name the following salts according to the old *-ic*, *-ous* system and the Stock system.

- a. CrS g. HgCl_2 m. CoO
 b. Cr_2O_3 h. Hg_2Cl_2 n. $\text{CoBr}_2 \cdot 6\text{H}_2\text{O}$
 c. $\text{CrI}_3 \cdot 6\text{H}_2\text{O}$ i. HgO o. SnF_4
 d. CuCl j. Fe_2O_3 p. SnO_2
 e. CuI_2 k. FeS q. Cu_2O
 f. $\text{CuBr}_2 \cdot 4\text{H}_2\text{O}$ l. $\text{FeI}_3 \cdot 6\text{H}_2\text{O}$ r. $\text{Fe}(\text{OH})_3$

Exploring Chemistry 1 edition Lab Manual Answers

WJ Hussar



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- Chapter 4: Exploring Chemistry 1edition Lab Manual Answers in Specific Contexts
- Chapter 5: Conclusion

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Table of Contents Exploring Chemistry 1edition Lab Manual Answers

1. Understanding the eBook Exploring Chemistry 1edition Lab Manual Answers
 - The Rise of Digital Reading Exploring Chemistry 1edition Lab Manual Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Exploring Chemistry 1edition Lab Manual Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Exploring Chemistry 1edition Lab Manual Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Exploring Chemistry 1edition Lab Manual Answers
 - Personalized Recommendations
 - Exploring Chemistry 1edition Lab Manual Answers User Reviews and Ratings
 - Exploring Chemistry 1edition Lab Manual Answers and Bestseller Lists
5. Accessing Exploring Chemistry 1edition Lab Manual Answers Free and Paid eBooks
 - Exploring Chemistry 1edition Lab Manual Answers Public Domain eBooks
 - Exploring Chemistry 1edition Lab Manual Answers eBook Subscription Services
 - Exploring Chemistry 1edition Lab Manual Answers Budget-Friendly Options
6. Navigating Exploring Chemistry 1edition Lab Manual Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Exploring Chemistry 1edition Lab Manual Answers Compatibility with Devices
 - Exploring Chemistry 1edition Lab Manual Answers Enhanced eBook Features
7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Exploring Chemistry 1edition Lab Manual Answers
 - Highlighting and Note-Taking Exploring Chemistry 1edition Lab Manual Answers
 - Interactive Elements Exploring Chemistry 1edition Lab Manual Answers
8. Staying Engaged with Exploring Chemistry 1edition Lab Manual Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Exploring Chemistry 1edition Lab Manual Answers
 9. Balancing eBooks and Physical Books Exploring Chemistry 1edition Lab Manual Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Exploring Chemistry 1edition Lab Manual Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Exploring Chemistry 1edition Lab Manual Answers
 - Setting Reading Goals Exploring Chemistry 1edition Lab Manual Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Exploring Chemistry 1edition Lab Manual Answers
 - Fact-Checking eBook Content of Exploring Chemistry 1edition Lab Manual Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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