

How are electrons shared to create covalently bonded molecules?

Why? Covalent bonds result from a sharing of electrons by two or more atoms (usually nonmetals). **Lewis theory** (Gilbert Newton Lewis, 1875-1946) focuses on the valence electrons, since the outermost electrons are the ones that are highest in energy and farthest from the nucleus, and are therefore the ones that are most exposed to other atoms when bonds form. **Lewis dot diagrams** for elements are a handy way of picturing valence electrons, and especially, what electrons are available to be shared in covalent bonds.

[illegible]

- Write the long hand electron configuration for H, Li, Na and circle the valence electrons.
 - How many valence electrons do the atoms H, Li and Na each have?
 - Does this agree with Model 1 showing the Lewis Dot Diagrams?
- How many valence electrons do the atoms B and Al each have (refer to Model 1)?

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