



Rates of Reaction Worksheet

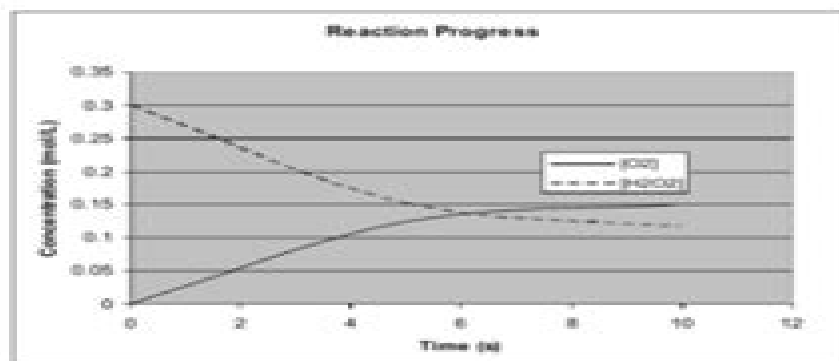
- What are the five most common ways to determine reaction rates?
- Solid phosphorus and oxygen gas react to form tetraphosphorus decoxide. Determine the average rate of reaction for oxygen during the first 40 s if the concentration changes from 0.200 mol/L to 0.000 mol/L during this time interval.
- At 40°C, hydrogen chloride gas will form from the reaction of gaseous hydrogen and chlorine, according to the following balanced chemical equation: $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow 2\text{HCl}(\text{g})$. Using the data provided, calculate the following average rates:

Time (s)	Concentration (mol/L)		
	$\text{H}_2(\text{g})$	$\text{Cl}_2(\text{g})$	$\text{HCl}(\text{g})$
0	1.000	1.000	0.000
2.16	0.500	0.500	1.000
4.32	0.250	0.250	1.500

- disappearance of hydrogen gas in the first 2.16 s. (0.231 mol/L.s)
- appearance of hydrogen chloride gas in the first 4.32 s (0.347 mol/L.s)
- disappearance of chlorine gas between 2.16 s and 4.32 s (0.116 mol/L.s)
- disappearance of hydrogen gas in the first 4.32 s (0.174 mol/L.s)

- Hydrogen peroxide in aqueous solution will decompose to produce oxygen gas and water. Use the graph to the right to:

- calculate the average reaction rate for the appearance of oxygen gas between 5 s and 10 s.
- calculate the average reaction rate of disappearance of hydrogen peroxide between 0 s and 5 s.
- calculate the instantaneous rate of appearance of oxygen gas at 4 s.



- Ammonia gas and oxygen gas react to form nitric oxide and water vapour. The instantaneous rate of consumption of ammonia is 2.0×10^{-2} mol/L at 10 s.
 - Write and balance the chemical equation.
 - Determine the instantaneous rate of oxygen gas consumption at 10 s. (2.2×10^{-2} mol/L.s)
 - Determine the instantaneous rate of water vapour appearance at 10 s. (2.7×10^{-2} mol/L.s)
- The chemical reaction represented by the following balanced chemical equation takes place in aqueous solution: $\text{MnO}_4^- (\text{aq}) + 5\text{Fe}^{2+} (\text{aq}) + 8\text{H}^+ (\text{aq}) \rightarrow \text{Mn}^{2+} (\text{aq}) + 5\text{Fe}^{3+} (\text{aq}) + 4\text{H}_2\text{O} (\text{l})$. When the disappearance of the permanganate ion, $\text{MnO}_4^- (\text{aq})$, is monitored, the rate of reaction for the consumption of the ion is 4.0×10^{-2} mol/L.s. Calculate the reaction rate for all other reactants and products. (Fe^{2+} : 0.20 mol/L.s, H^+ : 0.32 mol/L.s, Mn^{2+} : 0.040 mol/L.s, Fe^{3+} : 0.20 mol/L.s, H_2O : 0.16 mol/L.s)
- One mole of aqueous sulfuric acid reacts with 2 mol of aqueous sodium hydrogen carbonate in an acid-base reaction.
 - Write a balanced chemical equation.
 - If 0.038 mol sodium hydrogen carbonate reacts in 20 s, calculate the average reaction rate as:
 - Grams of sodium hydrogen carbonate consumed per second. (0.16 g/s)
 - Moles of sulfuric acid consumed per second. (9.5×10^{-4} mol/s)
 - Moles of carbon dioxide gas formed per second. (1.9×10^{-3} mol/s)

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