

# 7 Ionic and Metallic Bonding

## **Big idea** BONDING AND INTERACTIONS

### 7.1 Ions

For students using the Foundation edition, assign problems 1, 3–5, 7–12, 14, 15, 18–20.

**Essential Understanding** Ions form when atoms gain or lose valence electrons, becoming electrically charged.

#### Lesson Summary

**Valence Electrons** Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation.

- ▶ For a representative element, the group number equals the number of valence electrons the atom contains.
- ▶ An electron dot structure shows the symbol of the element and its valence electrons.
- ▶ Atoms tend to gain or lose the number of electrons that will provide the atom with a noble gas electron configuration.

**Formation of Cations** Cations are positively charged ions formed when an atom loses one or more valence electrons.

- ▶ Atoms and the cations formed from them have different properties.
- ▶ Elements in Group 1 form cations with a charge of  $1+$ , and those in Group 2 form cations with a charge of  $2+$ .
- ▶ Many transition metals form more than one cation and do not follow the octet rule.

**Formation of Anions** Anions are negatively charged ions formed when an atom gains one or more valence electrons.

- ▶ Commonly, the name of an anion ends in *-ide*.
- ▶ Anions form from nonmetallic elements.
- ▶ The anions formed from halogens are known as halides.

After reading Lesson 7.1, answer the following questions.

#### Valence Electrons

1. What are valence electrons?

*Valence electrons are the electrons in the highest occupied energy level of an element's atoms.*

2. The valence electrons largely determine the *chemical properties* of an element and are usually the only electrons used in *chemical bonds*.

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**Elizabeth DePoy, Laura N. Gitlin**



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### **Table of Contents Lesson 13 1 Reading And Study Workbook Answers**

1. Understanding the eBook Lesson 13 1 Reading And Study Workbook Answers
  - The Rise of Digital Reading Lesson 13 1 Reading And Study Workbook Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Lesson 13 1 Reading And Study Workbook Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Lesson 13 1 Reading And Study Workbook Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson 13 1 Reading And Study Workbook Answers
  - Personalized Recommendations



- Lesson 13 1 Reading And Study Workbook Answers User Reviews and Ratings
- Lesson 13 1 Reading And Study Workbook Answers and Bestseller Lists
- 5. Accessing Lesson 13 1 Reading And Study Workbook Answers Free and Paid eBooks
  - Lesson 13 1 Reading And Study Workbook Answers Public Domain eBooks
  - Lesson 13 1 Reading And Study Workbook Answers eBook Subscription Services
  - Lesson 13 1 Reading And Study Workbook Answers Budget-Friendly Options
- 6. Navigating Lesson 13 1 Reading And Study Workbook Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Lesson 13 1 Reading And Study Workbook Answers Compatibility with Devices
  - Lesson 13 1 Reading And Study Workbook Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Lesson 13 1 Reading And Study Workbook Answers
  - Highlighting and Note-Taking Lesson 13 1 Reading And Study Workbook Answers
  - Interactive Elements Lesson 13 1 Reading And Study Workbook Answers
- 8. Staying Engaged with Lesson 13 1 Reading And Study Workbook Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Lesson 13 1 Reading And Study Workbook Answers
- 9. Balancing eBooks and Physical Books Lesson 13 1 Reading And Study Workbook Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Lesson 13 1 Reading And Study Workbook Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Lesson 13 1 Reading And Study Workbook Answers
  - Setting Reading Goals Lesson 13 1 Reading And Study Workbook Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lesson 13 1 Reading And Study Workbook Answers
  - Fact-Checking eBook Content of Lesson 13 1 Reading And Study Workbook Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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