

**Assignment #2 – Compound Names and Formulas**  
**Multi-valent ions only**

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide CuO
2. mercury (I) oxide Hg<sub>2</sub>O
3. gold (III) chloride AuCl<sub>3</sub>
4. thallium (III) bromide TlBr<sub>3</sub>
5. bismuth (V) oxide Bi<sub>2</sub>O<sub>5</sub>
6. terbium (III) oxide Tb<sub>2</sub>O<sub>3</sub>
7. uranium (VI) oxide UO<sub>3</sub>
8. protactinium (IV) bromide PaBr<sub>4</sub>
9. cerium (III) oxide Ce<sub>2</sub>O<sub>3</sub>
10. arsenic (V) sulphide As<sub>2</sub>S<sub>5</sub>

11. manganese (III) oxide Mn<sub>2</sub>O<sub>3</sub>
12. vanadium (II) bromide VBr<sub>2</sub>
13. niobium(V) oxide Nb<sub>2</sub>O<sub>5</sub>
14. titanium (III) oxide Ti<sub>2</sub>O<sub>3</sub>
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co<sub>3</sub>P<sub>2</sub>
18. tin (II) oxide SnO
19. thulium (II) bromide TmBr<sub>2</sub>
20. copper (II) bromide CuBr<sub>2</sub>

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. SnCl<sub>4</sub> Tin (IV) Chloride
2. BiBr<sub>5</sub> Bismuth (V) Bromide
3. PoO<sub>2</sub> Polonium (IV) Oxide
4. PbI<sub>2</sub> Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au<sub>2</sub>O<sub>3</sub> Gold (III) Oxide
8. FeCl<sub>2</sub> Iron (II) Chloride

9. PdF<sub>4</sub> Palladium (IV) Fluoride
10. Os<sub>2</sub>O<sub>3</sub> Osmium (III) Oxide
11. MoBr<sub>2</sub> Molybdenum (II) Bromide
12. VCl<sub>5</sub> Vanadium (V) Chloride
13. Mn<sub>2</sub>O<sub>3</sub> Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np<sub>2</sub>O<sub>3</sub> Neptunium (III) Oxide
16. V<sub>2</sub>O<sub>5</sub> Vanadium (V) Oxide

# Ionic Compounds Homework Review Sheet Answer Key

**R Bogdan**



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*Merrill Chemistry* Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 Chemical Principles Peter Atkins,Loretta Jones,2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry s frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text Instructors have the option of whether to incorporate calculus in the coverage of topics The multimedia integration of Chemical Principles is more deeply established than any other text for this course Through the unique eBook the comprehensive Chemistry Portal Living Graph icons that connect the text to the Web and a complete set of animations students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding

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