

A chemical bond - forms when 2 or more atoms rearrange valence
electrons to increase stability.

ionic bond - forms when valence electrons are transferred
from one atom to another

cation - atom loses electrons to become + charged

anion - atom gains electrons to become - charged

In ionic compounds the ions are arranged in a crystal lattice.
Strong forces hold the ions together.

Properties of ionic compounds:

- high melting and boiling points
- hard - not easily crushed
- conduct electricity when melted or
dissolved because the ions are free to move.

covalent bond - electrons are shared, forming
molecules.

Covalent compounds have weaker forces holding the
molecules together.

Properties of covalent compounds:

- lower melting and boiling points
- Many covalent compounds are volatile liquids or gases.
- soft - easier to crush
- are not conductors of electricity

electronegativity - property that tells how strong an atom's attraction
is for electrons.

Since oxygen has a higher electronegativity than hydrogen, oxygen
holds onto shared electrons tighter, giving the oxygen a partial
negative charge and the hydrogen a partial positive charge.

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PT Brinkman

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the PT Brinkman text.

Ionic Bonds And Notetaking Guide Answer:

Unveiling the Energy of Verbal Art: An Mental Sojourn through **Ionic Bonds And Notetaking Guide Answer**

In a world inundated with monitors and the cacophony of instantaneous communication, the profound power and mental resonance of verbal artistry usually fade in to obscurity, eclipsed by the constant barrage of noise and distractions. However, located within the lyrical pages of **Ionic Bonds And Notetaking Guide Answer**, a fascinating perform of fictional brilliance that impulses with raw feelings, lies an wonderful journey waiting to be embarked upon. Composed by a virtuoso wordsmith, this magical opus books readers on a mental odyssey, gently exposing the latent potential and profound impact stuck within the complex web of language. Within the heart-wrenching expanse of the evocative analysis, we will embark upon an introspective exploration of the book is key styles, dissect their fascinating writing type, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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