

- Directions:**
 1. An element may be assigned a positive or negative oxidation number of zero.
 2. An atom with a higher electronegativity is assigned a negative oxidation number.
 3. The sum of all oxidation numbers for a molecule must be zero.

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1. Write an equation that shows how the element is oxidized or reduced.
 2. Write the oxidation number next to the element.
 3. Write the oxidation number next to the element.

Complete the following table, using the periodic table in the back of your book.

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	NUMBER OF ELECTRONS LOST OR GAINED	TYPE OF ION
1	Sodium	Na^+	11	10	lost one	cation
2	Chlorine	Cl^-	17	18	gained one	anion
3	Sulfur	S^{2-}	16	18	gained two	anion
4	Indium	In^{3+}	49	46	lost three	cation
5	Calcium	Ca^{2+}	20	18	lost two	cation
6	Strontium	Sr^{2+}	38	36	lost two	cation
7	Barium	Ba^{2+}	56	54	lost two	cation
8	Hydrogen	H^+	1	0	lost one	cation
9	Lithium	Li^+	3	2	lost one	cation
10	Magnesium	Mg^{2+}	12	10	lost two	cation
11	Aluminum	Al^{3+}	13	10	lost three	cation
12	Gallium	Ga^{3+}	31	28	lost three	cation
13	Indium	In^+	49	48	lost one	cation
14	Thallium	Tl^+	81	80	lost one	cation
15	Lead	Pb^{2+}	82	80	lost two	cation
16	Carbon	C^{4-}	6	10	gained four	anion

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Xiang Xie

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