#### Acid and Base Worksheet - Answers

- Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair.
  - a) HNO₂ + OH¹ → H₂O + NO₃¹ HNO₂ and NO₃¹ make one pair OH¹ and H₂O make the other
  - b) CH<sub>3</sub>NH<sub>2</sub> + H<sub>2</sub>O → CH<sub>3</sub>NH<sub>3</sub>\* + OH<sup>1</sup>
     CH<sub>3</sub>NH<sub>2</sub> and CH<sub>3</sub>NH<sub>3</sub>\* make one pair OH<sup>1</sup> and H<sub>2</sub>O make the other
  - c) OH¹ + HPO₄² → H₂O + PO₄³ HPO₄² and PO₄³ make one pair OH¹ and H₂O make the other
- 2) The compound NaOH is a base by all three of the theories we discussed in class. However, each of the three theories describes what a base is in different terms. Use your knowledge of these three theories to describe NaOH as an Arrhenius base, a Brønsted-Lowry base, and a Lewis base.
  - NaOH is an Arrhenius base because it creates OH<sup>1</sup> ions when placed in water.
  - NaOH is a Brønsted-Lowry base because it accepts H<sup>\*1</sup> ions from acids.
  - NaOH is a Lewis base because the lone pairs on the hydroxide ion can be donated to other compounds.
- When hydrogen chloride reacts with ammonia, ammonium chloride is formed. Write the equation for this process, and indicate which of the reagents is the Lewis acid and which is the Lewis base.
  HCl + NH₃ → NH₄Cl
  NH₃ is a Lewis base because it uses its lone pair electrons to pull a hydrogen atom from hydrochloric acid.
  HCl is a Lewis acid because it accepts electrons from NH₃ when the H is transferred.
- 4) Write an equation for each of the following reactions and indicate what each reaction produces: (For each of the following oxides, indicate if it is an acid or basic anhydride).
  - Carbon dioxide and water
     CO₂ + H₂O → H₂CO₃
     Acid anhydride
  - b. Sodium oxide and water
     Na₂O + H₂O → 2NaOH(aq)
     Basic anhydride
  - c. Dichlorine trioxide and water
     Cl₂O₂ + H₂O → 2HClO₂
     Acidic anhydride
  - d. Beryllium oxide and water
     BeO + H₂O → Be(OH)₂
     Basic anhydride

# **Introduction To Acids And Bases Answer Key**

Neale Taylor, Angela Stubbs, Robert Stokes

#### **Introduction To Acids And Bases Answer Key:**

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Introduction to Porous Materials Pascal Van Der Voort, Karen Leus, Els De Canck, 2019-06-11 The first comprehensive textbook on the timely and rapidly developing topic of inorganic porous materials. This is the first textbook to completely cover a broad range of inorganic porous materials. It introduces the reader to the development of functional porous inorganic materials from the synthetic zeolites in the 50 s to today s hybrid materials such as metal organic frameworks MOFs covalent organic frameworks COFs and related networks. It also provides the necessary background to understand how porous materials are organized characterized and applied in adsorption catalysis and many other domains Additionally the book explains characterization and application from the materials scientist viewpoint giving the reader a practical approach on the characterization and application of the respective materials. Introduction to Inorganic Porous Materials begins by describing the basic concepts of porosity and the different types of pores surfaces and amorphous versus crystalline materials before introducing readers to nature s porous materials. It then goes on to cover everything from adsorption and catalysis to amorphous materials such as silica to inorganic carbons and Periodic Mesoporous Organosilicas PMOs It discusses the synthesis and applications of MOFs and the broad family of COFs It concludes with a look at future prospects and emerging

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