

- b) If the rate of consumption of magnesium is  $5.0 \times 10^{-9}$  mol/s, find the rate of consumption of HCl in moles/s.

$$5.0 \times 10^{-9} \frac{\text{mol Mg}}{\text{s}} \times \frac{2 \text{ mol HCl}}{1 \text{ mol Mg}} = 1.0 \times 10^{-8} \text{ mol HCl/s}$$

Answer  $1.0 \times 10^{-8} \text{ mol HCl/s}$

- c) If the rate of consumption of magnesium is  $5.0 \times 10^{-9}$  mol/s, find the rate of production of  $\text{H}_2$  in g/s.

$$5.0 \times 10^{-9} \frac{\text{mol Mg}}{\text{s}} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} =$$

Answer  $1.0 \times 10^{-8} \text{ g H}_2/\text{s}$

- d) If the rate of consumption of magnesium is  $5.0 \times 10^{-9}$  mol/s, find the rate of production of  $\text{H}_2$  in L/s (@STP).

$$5.0 \times 10^{-9} \frac{\text{mol Mg}}{\text{s}} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} =$$

Answer  $1.1 \times 10^{-7} \text{ L H}_2/\text{s}$

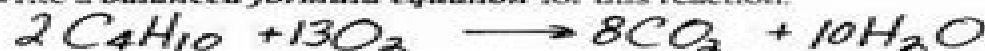
- e) If the rate of consumption of magnesium is  $5.0 \times 10^{-9}$  mol/s, find the mass of Mg consumed in 5.0 minutes.

$$5.0 \times 10^{-9} \frac{\text{mol Mg}}{\text{s}} \times \frac{24.3 \text{ g Mg}}{1 \text{ mol Mg}} \times \frac{60}{1} \times 5.0 \text{ min} = 3.6 \times 10^{-5} \text{ g Mg}$$

Answer  $3.6 \times 10^{-5} \text{ g Mg}$

3. When butane ( $\text{C}_4\text{H}_{10}$ ) is burned in air (oxygen), the products carbon dioxide and water are formed.

- a) Write a balanced formula equation for this reaction.



- b) If butane is consumed at an average rate of 0.116 grams/s, determine the rate of production of  $\text{CO}_2$  in g/s.

$$0.116 \frac{\text{g C}_4\text{H}_{10}}{\text{s}} \times \frac{1 \text{ mol C}_4\text{H}_{10}}{58.0 \text{ g C}_4\text{H}_{10}} \times \frac{8 \text{ mol CO}_2}{2 \text{ mol C}_4\text{H}_{10}} \times \frac{44.0 \text{ g CO}_2}{1 \text{ mol CO}_2} = 0.352 \text{ g CO}_2/\text{s}$$

Answer  $0.352 \text{ g CO}_2/\text{s}$

# Heath Chemistry Learning Guide Molar Relationships Answer

**Irene Cesa**



## Heath Chemistry Learning Guide Molar Relationships Answer:

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**Cumulated Index to the Books** ,1933      Heath Chemistry ,1993      **Study Guide for Chemical Principles [by] Steven S. Zumdahl** Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples      **Study Guide for Introductory Chemistry** Iris Stovall, 1990      **The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises      Solutions Guide for Introductory Chemistry James F. Hall, 1996      *Study Guide* Darrell D. Ebbing, Susan M. Schelble, 1998      Study Guide for Chemistry by Steven S. Zumdahl Martha B. Barrett, 1986      **Molar Relationships and Stoichiometry** Irene Cesa, 2002-01-01 Analyze the formulas of compounds and determine molar relationships among reactants and products      **Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble, 1994      **Introductory Chemistry Study Guide** Holder, 2003-05

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