



Part III: Balance the following chemical equations

20. $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{AlCl}_3$
21. $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{FeO} \rightarrow \underline{\hspace{1cm}} \text{Al}_2\text{O}_3 + \underline{\hspace{1cm}} \text{Fe}$
22. $\underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{NaCl}$
23. $\underline{\hspace{1cm}} \text{P}_4\text{O}_{10} + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{H}_3\text{PO}_4$
24. $\underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3 + \underline{\hspace{1cm}} \text{Ca}(\text{OH})_2 \rightarrow \underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{CaSO}_4$
25. $\underline{\hspace{1cm}} \text{C}_3\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
26. $\underline{\hspace{1cm}} \text{Fe}_2\text{O}_3 + \underline{\hspace{1cm}} \text{H}_2 \rightarrow \underline{\hspace{1cm}} \text{Fe} + \underline{\hspace{1cm}} \text{H}_2\text{O}$
27. $\underline{\hspace{1cm}} \text{N}_2\text{O}_5 + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{HNO}_3$
28. $\underline{\hspace{1cm}} \text{H}_3\text{AsO}_4 \rightarrow \underline{\hspace{1cm}} \text{As}_2\text{O}_5 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
29. $\underline{\hspace{1cm}} \text{Pb} + \underline{\hspace{1cm}} \text{Na} + \underline{\hspace{1cm}} \text{C}_2\text{H}_5\text{Cl} \rightarrow \underline{\hspace{1cm}} \text{Pb}(\text{C}_2\text{H}_5)_4 + \underline{\hspace{1cm}} \text{NaCl}$

Part IV: In each of the following problems you are given some potential reaction. Determine what type of reaction would occur, predict the products, write the complete chemical equation and balance it!

30. Aluminum powder burns in the presence of oxygen to form ...
31. Magnesium metal reacts with nitric acid to form ...
32. Aqueous barium chloride reacts with sodium hydroxide to form ...
33. Butane burns in the presence of oxygen to form ...
34. Hydrogen gas reacts with chlorine gas to form ...
35. Aluminum powder reacts with iron(III) oxide to form ...

Formulas And Chemical Reactions Study Guide

Julia Schneider



Formulas And Chemical Reactions Study Guide:

Formulas And Chemical Reactions Study Guide Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Formulas And Chemical Reactions Study Guide**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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