

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Exploring Periodic Trends Answer Key

**Jeylan T. Mortimer, Michael J.
Shanahan**



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Table of Contents Exploring Periodic Trends Answer Key

1. Understanding the eBook Exploring Periodic Trends Answer Key
 - The Rise of Digital Reading Exploring Periodic Trends Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Exploring Periodic Trends Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Exploring Periodic Trends Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Exploring Periodic Trends Answer Key
 - Personalized Recommendations
 - Exploring Periodic Trends Answer Key User Reviews and Ratings
 - Exploring Periodic Trends Answer Key and Bestseller Lists

5. Accessing Exploring Periodic Trends Answer Key Free and Paid eBooks
 - Exploring Periodic Trends Answer Key Public Domain eBooks
 - Exploring Periodic Trends Answer Key eBook Subscription Services
 - Exploring Periodic Trends Answer Key Budget-Friendly Options
6. Navigating Exploring Periodic Trends Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Exploring Periodic Trends Answer Key Compatibility with Devices
 - Exploring Periodic Trends Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Exploring Periodic Trends Answer Key
 - Highlighting and Note-Taking Exploring Periodic Trends Answer Key
 - Interactive Elements Exploring Periodic Trends Answer Key
8. Staying Engaged with Exploring Periodic Trends Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Exploring Periodic Trends Answer Key
9. Balancing eBooks and Physical Books Exploring Periodic Trends Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Exploring Periodic Trends Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Exploring Periodic Trends Answer Key
 - Setting Reading Goals Exploring Periodic Trends Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Exploring Periodic Trends Answer Key
 - Fact-Checking eBook Content of Exploring Periodic Trends Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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