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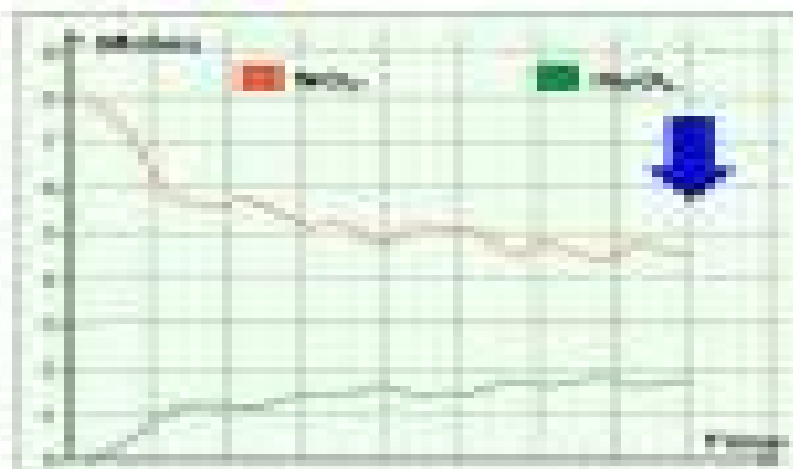
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### Equilibrium & Le Chatelier's

A closed flask containing hydrogen gas at high pressure (100 atm) and some solid hydrogen chloride (solid) is placed in a water bath (25°C). The graph below shows the hydrogen chloride concentration (molar) versus time (min) after the flask is placed in the water bath.



- The concentration of the hydrogen gas,  $H_2(g)$ , increases.
- The reaction between hydrogen and hydrogen chloride is faster.
- The rate of hydrogen chloride formation is constant; the rate of hydrogen gas formation is constant.
- The activation energy of the reaction does not change.

**Correct Answer:** B. Because the concentration of hydrogen gas is the same as the concentration of hydrogen chloride.

Hydrogen gas and hydrogen chloride are in a closed flask. The reaction is  $H_2(g) + HCl(g) \rightleftharpoons H_2(g) + HCl(g)$ . The graph below shows the concentration of hydrogen gas versus time. The graph is labeled with the concentration of hydrogen gas and the concentration of hydrogen chloride.

# Explorelearning Equilibrium And Concentration Gizmo

**Yi-Tong Ma**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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