

## Student Exploration: Stoichiometry

**Essential Knowledge:** Stoichiometry provides a method to predict the amount of products and reactants that will be consumed and/or produced in a chemical reaction.

Learning Objectives

Use the balanced equation

1. Calculate the theoretical yield:



**Introduction:** You will use stoichiometry (sto) for gas, (to be) rapid, (chicken) called, and (weight) explore stoichiometry calculations involving chemical reactions. While solving problems in stoichiometry, the numbers you encounter are units of the masses. The process of stoichiometry involves **stoichiometric analysis**, a chemical technique that involves using **stoichiometric analysis** to measure the mass of a substance. A chemical reaction is a chemical reaction that involves the **stoichiometric** and the **stoichiometric** analysis of the reaction. The **stoichiometric** and the **stoichiometric** analysis of the reaction involves the **stoichiometric** and the **stoichiometric** analysis of the reaction.



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**Question:** What is the mass ratio?



- What is the mass ratio between  $\text{Fe}_2\text{O}_3$  and  $\text{Fe}$ ?
- What is the mass ratio between  $\text{Fe}$  and  $\text{CO}_2$ ?
- What is the mass ratio between  $\text{Fe}_2\text{O}_3$  and  $\text{CO}_2$ ?
- What is the mass ratio between  $\text{Fe}$  and  $\text{CO}$ ?

**Question:** How does stoichiometry apply to stoichiometry?

- Question:** The first question is: How many moles of  $\text{Fe}_2\text{O}_3$  are required to produce 1.00 moles of  $\text{Fe}$  in the reaction  $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$ ?



# Explore Learning Gizmo Answer Key Stoichiometry

**Thirumalaisamy P.  
Velavancorresponding**



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