

ACID-BASE RELATIVE STRENGTHS

in Aqueous Solutions at 25° C

Name of Acid	Acid		Base		K _a
Hydriodic	HI	→	H ⁺ + I ⁻		>> 1
Hydrobromic	HBr	→	H ⁺ + Br ⁻		>> 1
Hydrochloric	HCl	→	H ⁺ + Cl ⁻		>> 1
Nitric	HNO ₃	→	H ⁺ + NO ₃ ⁻		> 1
Sulfuric	H ₂ SO ₄	→	H ⁺ + HSO ₄ ⁻		> 1
Oxalic	H ₂ C ₂ O ₄	⇌	H ⁺ + HC ₂ O ₄ ⁻		5.90 × 10 ⁻²
Sulfurous	H ₂ SO ₃	⇌	H ⁺ + HSO ₃ ⁻		1.54 × 10 ⁻²
Hydrogen Sulfate Ion	HSO ₄ ⁻	⇌	H ⁺ + SO ₄ ²⁻		1.20 × 10 ⁻²
Phosphoric	H ₃ PO ₄	⇌	H ⁺ + H ₂ PO ₄ ⁻		7.52 × 10 ⁻³
Nitrous	HNO ₂	⇌	H ⁺ + NO ₂ ⁻		4.6 × 10 ⁻⁴
Hydrofluoric	HF	⇌	H ⁺ + F ⁻		3.53 × 10 ⁻⁴
Formic	HCO ₂ H	⇌	H ⁺ + HCO ₂ ⁻		1.77 × 10 ⁻⁴
Benzoic	C ₆ H ₅ CO ₂ H	⇌	H ⁺ + C ₆ H ₅ CO ₂ ⁻		6.46 × 10 ⁻⁵
Acetic	CH ₃ CO ₂ H	⇌	H ⁺ + CH ₃ CO ₂ ⁻		1.76 × 10 ⁻⁵
Carbonic	H ₂ CO ₃	⇌	H ⁺ + HCO ₃ ⁻		4.30 × 10 ⁻⁷
Hydrogen Sulfite Ion	HSO ₃ ⁻	⇌	H ⁺ + SO ₃ ²⁻		1.02 × 10 ⁻⁷
Dihydrogen Phosphate Ion	H ₂ PO ₄ ⁻	⇌	H ⁺ + HPO ₄ ²⁻		6.23 × 10 ⁻⁸
Boric	H ₃ BO ₃	⇌	H ⁺ + H ₂ BO ₃ ⁻		7.3 × 10 ⁻¹⁰
Ammonium Ion	NH ₄ ⁺	⇌	H ⁺ + NH ₄ ⁺		5.64 × 10 ⁻¹⁰
Hydrogen Carbonate Ion	HCO ₃ ⁻	⇌	H ⁺ + CO ₃ ²⁻		5.61 × 10 ⁻¹¹
Monohydrogen Phosphate Ion	HPO ₄ ²⁻	⇌	H ⁺ + PO ₄ ³⁻		2.2 × 10 ⁻¹³
Water	H ₂ O	⇌	H ⁺ + OH ⁻		1.0 × 10 ⁻¹⁴

For space considerations, all reactions omit water as a reactant



Lab 16 1 Creating Acid Base Strength Table

SA Adler



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