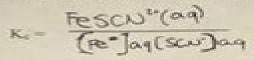
Report Sheet - Experiment 13-Chemical Equilibrium

- 1. Write the chemical equation for the reversible reaction studied in this experiment:

 Fe *(aq)+ SCN * > Fe (SCN) *(aq)
- 2. Write the equilibrium expression for the reaction studied in this experiment:



3. Complete the following table from your observations of test tubes #1-#6:

Test Tube	Change	Initial Color	Final Color (darker/lighter)
	Initial color before anything is added		
1	Added Fe(NO ₁) ₁	Reddish color	dorker
2	Added KSCN	Reddish color	darker
3	Added NH ₄ CI	Reddish color	Lighter
4	Added AgNO	Reddish color	Lighter
5	Cold	Reddish color	darker
6	Hot	Reddish color	lighter

4. Complete the following table for test tubes #2 - 4:

Test Tube	What concentration was changed and how it was changed?	The equilibrium shifted to left/right/no change	What evidence for the direction of shift do you have?
2	Fesch to a mix	Shifted Aykt	principle but react on from left to right accust was when more reactant add
3	[SCN] was	Shifted aight	as more reactant was added, the concentration of the product goes up
4	Fezz concentration was	Shifetell Just+	Fe COL WILL Form which leads to a lower Fe to concentration

Lab 1rate And Equilibrium Report Answer

Terry C. Jones

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Table of Contents Lab 1rate And Equilibrium Report Answer

- 1. Understanding the eBook Lab 1rate And Equilibrium Report Answer
 - The Rise of Digital Reading Lab 1rate And Equilibrium Report Answer
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Lab 1rate And Equilibrium Report Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab 1rate And Equilibrium Report Answer
 - User-Friendly Interface

- 4. Exploring eBook Recommendations from Lab 1rate And Equilibrium Report Answer
 - Personalized Recommendations
 - Lab 1rate And Equilibrium Report Answer User Reviews and Ratings
 - Lab 1rate And Equilibrium Report Answer and Bestseller Lists
- 5. Accessing Lab 1rate And Equilibrium Report Answer Free and Paid eBooks
 - Lab 1rate And Equilibrium Report Answer Public Domain eBooks
 - Lab 1rate And Equilibrium Report Answer eBook Subscription Services
 - Lab 1rate And Equilibrium Report Answer Budget-Friendly Options
- 6. Navigating Lab 1rate And Equilibrium Report Answer eBook Formats
 - o ePub, PDF, MOBI, and More
 - Lab 1rate And Equilibrium Report Answer Compatibility with Devices
 - Lab 1rate And Equilibrium Report Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab 1rate And Equilibrium Report Answer
 - Highlighting and Note-Taking Lab 1rate And Equilibrium Report Answer
 - Interactive Elements Lab 1rate And Equilibrium Report Answer
- 8. Staying Engaged with Lab 1rate And Equilibrium Report Answer
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab 1rate And Equilibrium Report Answer
- 9. Balancing eBooks and Physical Books Lab 1rate And Equilibrium Report Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab 1rate And Equilibrium Report Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lab 1rate And Equilibrium Report Answer
 - Setting Reading Goals Lab 1rate And Equilibrium Report Answer
 - Carving Out Dedicated Reading Time

- 12. Sourcing Reliable Information of Lab 1rate And Equilibrium Report Answer
 - Fact-Checking eBook Content of Lab 1rate And Equilibrium Report Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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