

Calculations need to be based on the limiting reactant.

- Example 2: What is the mass of sulfur dioxide that can be produced by the reaction of 95.6 g carbon disulfide with 100. g oxygen?



$$\frac{95.6\text{g CS}_2}{76.2\text{g CS}_2} \times \frac{1\text{mol CS}_2}{1\text{mol CS}_2} \times \frac{2\text{mol SO}_2}{1\text{mol CS}_2} \times \frac{64.1\text{g SO}_2}{1\text{mol SO}_2} = 161\text{g SO}_2$$

$$\frac{100.\text{g O}_2}{32.0\text{g O}_2} \times \frac{1\text{mol O}_2}{1\text{mol O}_2} \times \frac{2\text{mol SO}_2}{3\text{mol O}_2} \times \frac{64.1\text{g SO}_2}{1\text{mol SO}_2} = 134\text{g SO}_2$$

134 g of SO_2 can be produced in this reaction.

Holt Chemistry Limiting Reactants Answers

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Holt Chemistry Limiting Reactants Answers:

Unveiling the Energy of Verbal Artistry: An Emotional Sojourn through **Holt Chemistry Limiting Reactants Answers**

In some sort of inundated with monitors and the cacophony of quick connection, the profound energy and emotional resonance of verbal artistry often fade into obscurity, eclipsed by the regular onslaught of noise and distractions. However, located within the musical pages of **Holt Chemistry Limiting Reactants Answers**, a captivating work of fictional elegance that pulses with natural emotions, lies an unforgettable journey waiting to be embarked upon. Published by way of a virtuoso wordsmith, this exciting opus instructions viewers on a psychological odyssey, lightly revealing the latent potential and profound impact embedded within the elaborate web of language. Within the heart-wrenching expanse of the evocative analysis, we will embark upon an introspective exploration of the book is key themes, dissect their captivating publishing fashion, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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