

8 • Electron Configurations & Periodicity**WRITING ELECTRON CONFIGURATIONS**

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ s: <u>18</u>	Element: Mg # of e ⁻ s: <u>12</u>	Element: N # of e ⁻ s: <u>7</u>	Element: Li # of e ⁻ s: <u>3</u>	Element: P # of e ⁻ s: <u>15</u>	Element: Cl # of e ⁻ s: <u>17</u>

Write the electron configurations of each of these in long form and short form:

1. Ar $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6$

Ar $[\text{Ne}] \ 3s^2 \ 3p^6$

2. Mg $1s^2 \ 2s^2 \ 2p^6 \ 3s^2$

Mg $[\text{Ne}] \ 3s^2$

3. N $1s^2 \ 2s^2 \ 2p^3$

N $[\text{He}] \ 2s^2 \ 2p^3$

4. Li $1s^2 \ 2s^1$

Li $[\text{He}] \ 2s^1$

5. P $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^3$

P $[\text{Ne}] \ 3s^2 \ 3p^3$

6. Cl $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^5$

Cl $[\text{Ne}] \ 3s^2 \ 3p^5$



7. Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.

Fe $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 3d^6 \ 4s^2$

Fe $[\text{Ar}] \ 3d^6 \ 4s^2$

Electron Configuration Practice Orbital Answers

Joacim Rocklöv

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Table of Contents Electron Configuration Practice Orbital Answers

1. Understanding the eBook Electron Configuration Practice Orbital Answers
 - The Rise of Digital Reading Electron Configuration Practice Orbital Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Practice Orbital Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Practice Orbital Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Practice Orbital Answers
 - Personalized Recommendations
 - Electron Configuration Practice Orbital Answers User Reviews and Ratings

- Electron Configuration Practice Orbital Answers and Bestseller Lists
- 5. Accessing Electron Configuration Practice Orbital Answers Free and Paid eBooks
 - Electron Configuration Practice Orbital Answers Public Domain eBooks
 - Electron Configuration Practice Orbital Answers eBook Subscription Services
 - Electron Configuration Practice Orbital Answers Budget-Friendly Options
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 - ePUB, PDF, MOBI, and More
 - Electron Configuration Practice Orbital Answers Compatibility with Devices
 - Electron Configuration Practice Orbital Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Practice Orbital Answers
 - Highlighting and Note-Taking Electron Configuration Practice Orbital Answers
 - Interactive Elements Electron Configuration Practice Orbital Answers
- 8. Staying Engaged with Electron Configuration Practice Orbital Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Practice Orbital Answers
- 9. Balancing eBooks and Physical Books Electron Configuration Practice Orbital Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Practice Orbital Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Practice Orbital Answers
 - Setting Reading Goals Electron Configuration Practice Orbital Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Practice Orbital Answers
 - Fact-Checking eBook Content of Electron Configuration Practice Orbital Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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