

7&8 • Atomic Structure & Periodicity

PRACTICE TEST

$Rhc = A$

$\lambda A = 2.18 \times 10^{-18} \text{ J}$	$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$
$R = 1.097 \times 10^7 \text{ m}^{-1}$	$c = 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}$
mass of an electron = $9.11 \times 10^{-31} \text{ kg}$	

1. What wavelength corresponds to a frequency of $8.22 \times 10^9 \text{ Hz}$?

$\lambda = \frac{c}{\nu}$

- a) 0.307 m d) 0.110 m
 b) **0.0365 m** $\frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{8.22 \times 10^9 \text{ s}^{-1}} = 0.036496 \text{ m}$
 c) 0.122 m

2. A radio station transmits at 110 MHz ($110 \times 10^6 \text{ Hz}$). What wavelength is this radio wave?

- a) $3.65 \times 10^{-5} \text{ m}$ c) $3.81 \times 10^{-5} \text{ m}$
 b) 3.30 m **d) 2.73 m**
 $\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{110 \times 10^6 \text{ s}^{-1}} = 2.73 \text{ m}$

3. Which one of the following is NOT a proper unit for frequency?

- a) Hz **c) $\text{m}\cdot\text{s}^{-1}$ velocity**
 b) s^{-1} d) $\frac{1}{\text{sec}}$

4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.

$n=6 \rightarrow n=2$
 $E_6 = \frac{-2.18 \times 10^{-18} \text{ J}}{36} - E_2 = \frac{-2.18 \times 10^{-18} \text{ J}}{4}$

- a) 0.12334 m **d) $4.1029 \times 10^{-7} \text{ m}$**
 b) 24.373 m e) 36.559 m
 c) $2.7353 \times 10^{-7} \text{ m}$

$E = \frac{hc}{\lambda} \quad \lambda = \frac{hc}{E}$
 $\Delta E = 4.84 \times 10^{-19} \text{ J} \quad = 4.1 \times 10^{-7} \text{ m}$

5. What is the relationship between the energy of a photon of light and its frequency?

- a) $E = \nu$ d) $E = \frac{1}{h\nu}$
 b) $E = \frac{h}{\nu}$ e) $E = \frac{\nu}{h}$
c) $E = h\nu$

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?

$\Delta E = E_3 - E_2$

- a) $1.52 \times 10^{-19} \text{ J}$ d) $4.48 \times 10^{-19} \text{ J}$
 b) $3.63 \times 10^{-19} \text{ J}$ **e) $3.03 \times 10^{-19} \text{ J}$**
 c) $2.18 \times 10^{-19} \text{ J}$
 $\frac{-2.18 \times 10^{-18} \text{ J}}{9} - \frac{-2.18 \times 10^{-18} \text{ J}}{4}$

7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.

- a) $3.03 \times 10^{-12} \text{ m}$** c) $3.30 \times 10^{11} \text{ m}$
 b) $2.42 \times 10^{-12} \text{ m}$ d) $1.59 \times 10^{-25} \text{ m}$
 $\lambda = \frac{h}{m\nu} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{9.11 \times 10^{-31} \text{ kg} \cdot 0.8 \cdot 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}$

8. What resultant is expected from the interference of the two waves shown below?



- a) **c)**
 b) d)

9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?

- a) n c) m_l
b) l d) m_s

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