

Naming Molecular Compounds

How are the chemical formula and name of a molecular compound related?

Why?

When you began chemistry class this year, you probably already knew that the chemical formula for carbon dioxide was CO_2 . Today you will find out why CO_2 is named that way. Naming chemical compounds correctly is of paramount importance. The slight difference between the names carbon monoxide (CO , a poisonous, deadly gas) and carbon dioxide (CO_2 , a greenhouse gas that we exhale when we breathe out) can be the difference between life and death! In this activity you will learn the naming system for molecular compounds.

Model 1 – Molecular Compounds

Molecular Formula	Number of Atoms of First Element	Number of Atoms of Second Element	Name of Compound
ClF	1	1	Chlorine monofluoride
ClF_5	1	5	Chlorine pentafluoride
CO	1	1	Carbon monoxide
CO_2	1	2	Carbon dioxide
Cl_2O	2	1	Dichlorine monoxide
PCl_5	1	5	Phosphorus pentachloride
N_2O_5	2	5	Dinitrogen pentoxide

1. Fill in the table to indicate the number of atoms of each type in the molecular formula.

2. Examine the molecular formulas given in Model 1 for various molecular compounds.

a. How many different *elements* are present in each compound shown?

Each compound has two elements.

b. Do the compounds combine metals with metals, metals with nonmetals, or nonmetals with nonmetals?

The compounds have nonmetals with nonmetals.

c. Based on your answer to b, what type of bonding must be involved in molecular compounds?

A covalent bond must be involved.

3. Find all of the compounds in Model 1 that have chlorine and fluorine in them. Explain why the name "chlorine fluoride" is not sufficient to identify a specific compound.

ClF and ClF_5 : Chlorine fluoride does not specify that the two compounds have different ratios of chlorine and fluorine.

4. Assuming that the name of the compound gives a clue to its molecular formula, predict how many atoms each of these prefixes indicates, and provide two examples.

mono – *one* – CO (carbon monoxide) and Cl_2O (dichlorine monoxide)

di – *two* – CO_2 (carbon dioxide) and Cl_2O (dichlorine monoxide)

penta – *five* – ClF_5 (chlorine pentafluoride) and PCl_5 (phosphorus pentachloride)

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Michael Brown



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