

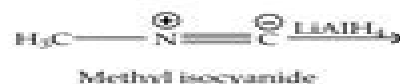
# SOLUTIONS

## Specimen Paper 2023

### Chemistry

#### Section-A

1. (A) (i) anisole, Williamson's synthesis.  
 (ii) three,  $s^{-1}$   
 (iii) lewis acid, lewis base.  
 (iv) acetone, iodoform.  
 (B) (i) Option (b) is correct.  
*Explanation:*



- (ii) Option (c) is correct.

*Explanation:* For a spontaneous reaction, the standard change in standard Gibbs free energy,  $\Delta G^\circ$  and the standard cell potential,  $E_{\text{cell}}^\circ$  will be negative and positive respectively.

- (iii) Option (c) is correct.

*Explanation:* Electronic configuration of  $\text{Cr} = [\text{Ar}] 3d^5 4s^1$  and  $\text{Cu} = [\text{Ar}] 3d^{10} 4s^1$

- (iv) Option (a) is correct.

*Explanation:* Since, for first order reaction, rate law is expressed as

$$k = \frac{2.303}{t} \log \frac{[R_0]}{[R]}$$

At  $t = 8 \text{ min}$ ,

$$k = \frac{2.303}{8} \log \frac{100}{25} = 0.1733 \text{ min}^{-1}$$

For 200 g,

$$k = \frac{2.303}{0.1733} \log \frac{200}{50} = 8 \text{ min}$$

- (C) (i) Phenol  
 (c) Neutral  $\text{FeCl}_3$  solution  
 (ii) Ethylenediamine  
 (d) Bidentate ligand  
 (iii) Colligative property  
 (a) Osmotic pressure  
 (iv) Amino acid  
 (b) Zwitter ion

- (D) (i) Option (a) is correct.

*Explanation:* Conductivity always decreases with the decrease in concentration for both weak and strong electrolytes as the number of ions per unit volume that carry the current

in a solution decreases on dilution.

- (ii) Option (d) is correct.

*Explanation:* Chlorination of nitrobenzene (not the nitration of chlorobenzene) leads to the formation of *m*-nitrochlorobenzene because  $-\text{NO}_2$  group deactivates the ring because it is meta directing.

#### Section-B

2. Osmotic pressure is a colligative property. According to it,

$$\pi = CRT$$

$$\pi = \frac{n}{V} RT \quad \left( C = \frac{n}{V} = \frac{m}{MV} \right)$$

$m$  = Mass of solute = 20 g

$M$  = Molar mass

$V$  = volume of solution = 500 mL

= 0.5 L

$R$  = Gas constant = 0.0821 atm

$\text{mol}^{-1} \text{K}^{-1}$

$T$  = Temperature = 298 K

Place the values in the above equation

$$\pi = \frac{m}{MV} RT$$

$$0.016 = \frac{20 \times 1000}{M \times 500} \times 0.0821 \times 298$$

$$M = \frac{20 \times 1000}{0.016 \times 500} \times 0.0821 \times 298$$

$$= 61164.5 \text{ g}$$

3. (i) Transition elements form a large number of complex compounds due to the comparatively smaller sizes of the metal ions and their high ionic charges and also the availability of

# Isc Chemistry Paper 24

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web figure 1 female injection site nipples umbilical cord figure 2 female genital papilla urogenital opening anus figure 3 male scrotum figure 4 male urogenital opening penis anus preparation and initial cuts tie one front leg of the animal with a string that passes underneath the dissecting pan to the other leg

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web obtain a fetal pig and identify the structures listed in figure 1 use figures 1 4 below to identify its sex use your pig and also a pig of the opposite sex to identify the structures in the photographs below the word urogenital refers to an opening that serves both the urinary excretory and the reproductive systems figure 1

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web 1 most of the pig s external features are familiar to you ears nose eyes etc on the belly you will see the umbilical cord which connected the fetal pig to its mother s placenta on either side of the umbilical cord you may see mammary papillae little nipples that will turn into teats in female pigs 2

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