

Chapter 9 Study Guide

Holt Chemistry: Stoichiometry

Questions

1. What is stoichiometry? Finding mass of product formed (or reactant consumed) when given a mass of another species in the reaction, by using the mole to mole ratio from the balanced equation.
2. A balanced equation verifies the law of conservation of mass: Total mass of reactants must equal total mass of products.
3. Give the 3 steps of converting the mass (grams) of one substance to the mass (grams) of another substance: 1. Convert grams to moles using molar mass; 2. Mole to Mole ratio from balanced equation; 3. Convert moles back to grams using molar mass.

Holt Chemistry Study Guide Concept Review

Stoichiometry

**John W. Moore, Conrad L.
Stanitski, Peter C. Jurs**



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books which supplements the AAAS science book list 3rd ed 1970 and the AAAS science book list supplement 1978 Films and Other Materials for Projection Library of Congress, 1974 General Chemistry James E. Brady, Gerard E. Humiston, 1982 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems

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