

Molarity Calculations

Calculate the molarities of the following solutions:



- 1) 2.3 moles of sodium chloride in 0.45 liters of solution.

$$M = \frac{2.3 \text{ moles}}{0.45 \text{ L}} = 5.111 \frac{\text{mol}}{\text{L}} = \boxed{5.11 \text{ M}}$$



- 2) 1.2 moles of calcium carbonate in 1.22 liters of solution.

$$M = \frac{1.2 \text{ mol}}{1.22 \text{ L}} = 0.983606557 \frac{\text{mol}}{\text{L}} = \boxed{0.98 \text{ M}}$$



- 3) 0.09 moles of sodium sulfate in 12 mL of solution.

12 mL = 0.012 L

$$M = \frac{0.09 \text{ mol}}{0.012 \text{ L}} = \boxed{7.5 \text{ M}}$$



- 4) 0.75 moles of lithium fluoride in 65 mL of solution.

65 mL = 0.065 L

$$M = \frac{0.75 \text{ mol}}{0.065 \text{ L}} = 11.53846154 \text{ M} = \boxed{11.54 \text{ M}}$$



- 5) 0.8 moles of magnesium acetate in 5 liters of solution.

$$M = \frac{0.8 \text{ mol}}{5 \text{ L}} = \boxed{0.16 \text{ M}}$$



- 6) 120 grams of calcium nitrite in 240 mL of solution.

240 mL = 0.240 L

$$n = 120 \text{ g} \left(\frac{1 \text{ mol}}{132.09 \text{ g}} \right)$$

MW_{Ca(NO₂)₂} = 132.09 g/mol

n = 0.91 mol

$$M = \frac{0.91 \text{ mol}}{0.240 \text{ L}} = \boxed{3.79 \text{ M}}$$



- 7) 98 grams of sodium hydroxide in 2.2 liters of solution.

MW_{NaOH} = 39.997 g/mol

$$n = 98 \text{ g} \left(\frac{1 \text{ mol}}{39.997 \text{ g}} \right) = 2.45 \text{ mol}$$

$$M = \frac{2.45 \text{ mol}}{2.2 \text{ L}} = \boxed{1.11 \text{ M}}$$



- 8) 1.2 grams of hydrochloric acid in 25 mL of solution.

25 mL = 0.025 L

MW_{HCl} = 36.461 g/mol

$$n = (1.2 \text{ g}) \left(\frac{1 \text{ mol}}{36.461 \text{ g}} \right) = 0.033 \text{ mol}$$

$$M = \frac{0.033 \text{ mol}}{0.025 \text{ L}} = \boxed{1.32 \text{ M}}$$



- 9) 45 grams of ammonia in 0.75 L of solution.

MW_{NH₃} = 17.031 g/mol

$$n = 45 \text{ g} \left(\frac{1 \text{ mol}}{17.031 \text{ g}} \right) = 2.64 \text{ mol}$$

$$M = \frac{2.64 \text{ mol}}{0.75 \text{ L}} = \boxed{3.52 \text{ M}}$$

General Chemistry Unit 11 Worksheet 4 Molar Concentration

Scott C. Dulebohn



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Table of Contents General Chemistry Unit 11 Worksheet 4 Molar Concentration

1. Understanding the eBook General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - The Rise of Digital Reading General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Personalized Recommendations
 - General Chemistry Unit 11 Worksheet 4 Molar Concentration User Reviews and Ratings
 - General Chemistry Unit 11 Worksheet 4 Molar Concentration and Bestseller Lists
5. Accessing General Chemistry Unit 11 Worksheet 4 Molar Concentration Free and Paid eBooks
 - General Chemistry Unit 11 Worksheet 4 Molar Concentration Public Domain eBooks
 - General Chemistry Unit 11 Worksheet 4 Molar Concentration eBook Subscription Services
 - General Chemistry Unit 11 Worksheet 4 Molar Concentration Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - General Chemistry Unit 11 Worksheet 4 Molar Concentration Compatibility with Devices
 - General Chemistry Unit 11 Worksheet 4 Molar Concentration Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Highlighting and Note-Taking General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Interactive Elements General Chemistry Unit 11 Worksheet 4 Molar Concentration
8. Staying Engaged with General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Unit 11 Worksheet 4 Molar Concentration
9. Balancing eBooks and Physical Books General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Unit 11 Worksheet 4 Molar Concentration
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Setting Reading Goals General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Fact-Checking eBook Content of General Chemistry Unit 11 Worksheet 4 Molar Concentration
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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