

4NB – VIDHATA

Date: 30/09/2014

DURATION : 2 Hrs 30 Min.

MAX MARKS : 75

INSTRUCTIONS TO THE STUDENTS :

- All questions are compulsory.
- Figures to the right indicate the marks.
- Draw neat sketches / diagrams / graphs / Flow sheets whenever necessary.
- Use of simple / scientific calculator / log tables permitted.

- A.] Answer Any Two of the following.

 - Predict the shape of following molecules :
 BeCl_2 , NH_3 , PCl_5 . 06
 - State and explain Ionization Energy. 03
 - What is Pauli's Rule? 03
 - Write short note on : Covalent Bond. 03
 - What are limitations of VSEPR theory? 03
 - What is Electronegativity? How does it vary in period and group in the periodic table? 03
 - Explain the Heisenberg's principle of Uncertainty? 03

B.] Answer Any Two of the following.

 - What is lone pair effect? How does it affect the bond angles in molecules? 04
 - What are applications of VSEPR theory? 04
 - Explain: BF_3 is planar whereas NH_3 is not planar. 04
 - Discuss the Rutherford's atomic model. 04
- A.] Answer Any Two of the following.

 - Define the following :-
Hydrogen Bond, Ionic Bond, Dipole Moment. 03
 - Calculate the formal charge on the atom in the following :
Nitrogen in Ammonium ion
Oxygen in Water
Nitrogen in Nitromethane 03
 - What are Dry Test and Wet Test? Give any two examples for each test. 06
 - What are types of Arrows and their significance? 03
 - Why Formic Acid is stronger than Acetic Acid? 03
 - Explain the structure of following functional groups
Ether, Cyanide, Ketone. 06

B.] Answer Any Two of the following.

 - What is Inductive Effect? Explain its types. 04
 - What is Semi-micro Analysis? 04
 - What is Formal Charge? Explain any two examples. 04
 - Explain sp^3 hybridization for Nitrogen and Oxygen. 04

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Hussin A.Rothana



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Table of Contents Fybsc October 2014

1. Understanding the eBook Fybsc October 2014
 - The Rise of Digital Reading Fybsc October 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Fybsc October 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Fybsc October 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Fybsc October 2014
 - Personalized Recommendations
 - Fybsc October 2014 User Reviews and Ratings
 - Fybsc October 2014 and Bestseller Lists
5. Accessing Fybsc October 2014 Free and Paid eBooks
 - Fybsc October 2014 Public Domain eBooks
 - Fybsc October 2014 eBook Subscription Services
 - Fybsc October 2014 Budget-Friendly Options

6. Navigating Fybsc October 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Fybsc October 2014 Compatibility with Devices
 - Fybsc October 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Fybsc October 2014
 - Highlighting and Note-Taking Fybsc October 2014
 - Interactive Elements Fybsc October 2014
8. Staying Engaged with Fybsc October 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Fybsc October 2014
9. Balancing eBooks and Physical Books Fybsc October 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Fybsc October 2014
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Fybsc October 2014
 - Setting Reading Goals Fybsc October 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Fybsc October 2014
 - Fact-Checking eBook Content of Fybsc October 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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