

**Concept-Development  
Practice Page**

**32-1**

**Coulomb's Law**

1. The diagram is of a hydrogen atom.

- Label the proton in the nucleus with a  $+$  sign and the orbital electron with a  $-$  sign.
- The electrical interaction between the nucleus and the orbital electron is a force of \_\_\_\_\_ (attractive) \_\_\_\_\_ (repulsive).
- According to Coulomb's law,

$$F = k \frac{q_1 q_2}{r^2}$$

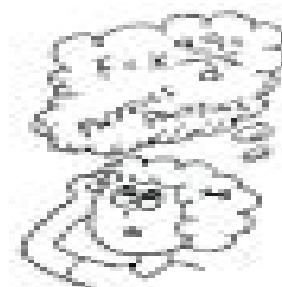
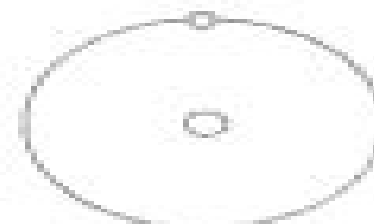
if the charge of either the nucleus or the orbital electron were greater, the force between the nucleus and the electron would be \_\_\_\_\_ (greater) \_\_\_\_\_ (less).

and if the distance between the nucleus and electron were greater the force would be \_\_\_\_\_ (greater) \_\_\_\_\_ (less).

If the distance between the nucleus and electron were doubled, the force would be \_\_\_\_\_ (1/4 as much) \_\_\_\_\_ (1/2 as much) \_\_\_\_\_ (two times as much) \_\_\_\_\_ (four times as much).

2. Consider the electric force between a pair of charged particles a certain distance apart. By Coulomb's law:

- If the charge on one of the particles is doubled, the force is \_\_\_\_\_ (unchanged) \_\_\_\_\_ (halved) \_\_\_\_\_ (doubled) \_\_\_\_\_ (quadrupled).
- If instead the charge on both particles is doubled, the force is \_\_\_\_\_ (unchanged) \_\_\_\_\_ (halved) \_\_\_\_\_ (doubled) \_\_\_\_\_ (quadrupled).
- If instead the distance between the particles is halved, the force is \_\_\_\_\_ (unchanged) \_\_\_\_\_ (halved) \_\_\_\_\_ (doubled) \_\_\_\_\_ (quadrupled).
- If the distance is halved and the charge of both particles is doubled, the force is \_\_\_\_\_ times as great.



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