

**Chapter 5 Chemistry: Electrons in Atoms****Matching***Match each item with the correct statement below.*

- |                              |                           |
|------------------------------|---------------------------|
| a. Pauli exclusion principle | c. electron configuration |
| b. aufbau principle          | d. ground state           |

- \_\_\_\_\_ 1. each orbital has at most two electrons  
 \_\_\_\_\_ 2. lowest energy level  
 \_\_\_\_\_ 3. tendency of electrons to enter orbitals of lowest energy first  
 \_\_\_\_\_ 4. arrangement of electrons around atomic nucleus

**Multiple Choice***Identify the choice that best completes the statement or answers the question.*

- \_\_\_\_\_ 5. What is the maximum number of electrons that can occupy the  $n = 3$  shell?  
 a. 50      b. 4      c. 8      d. 32      e. 18
- \_\_\_\_\_ 6. Which of the following elements is a d-block element?  
 a. Cd      b. Cf      c. C      d. Cs      e. Cl
- \_\_\_\_\_ 7. Which element has the electron configuration  $1s^2 2s^1 2p^6 3s^2 3p^6 3d^5 4s^2$ ?  
 a. Ag      b. Cd      c. Zn      d. Cu      e. Ge
- \_\_\_\_\_ 8. Which element has the following ground state electron configuration?  
 [Ar]   
 a. Fe      b. Co      c. Cu      d. Ni      e. Zn
- \_\_\_\_\_ 9. The  $n = \underline{\hspace{1cm}}$  shell is the lowest that may contain d-orbitals.  
 a. 2      b. 4      c. 3      d. 5      e. 6
- \_\_\_\_\_ 10. In Bohr's model of the atom, where are the electrons and protons located?  
 a. The electrons move around the protons, which are at the center of the atom.  
 b. The electrons and protons move throughout the atom.  
 c. The electrons occupy fixed positions around the protons, which are at the center of the atom.  
 d. The electrons and protons are located throughout the atom, but they are not free to move.
- \_\_\_\_\_ 11. How many energy sublevels are in the second principal energy level?  
 a. 1      c. 3  
 b. 2      d. 4
- \_\_\_\_\_ 12. What is the maximum number of orbitals in the p sublevel?  
 a. 2      c. 4  
 b. 3      d. 5

# Electrons In Atoms Chapter Test A

**Richard Bailey**

**Electrons In Atoms Chapter Test A:**

## Reviewing **Electrons In Atoms Chapter Test A**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Electrons In Atoms Chapter Test A**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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