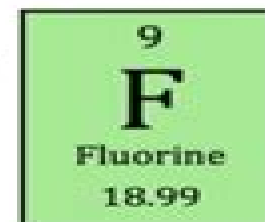
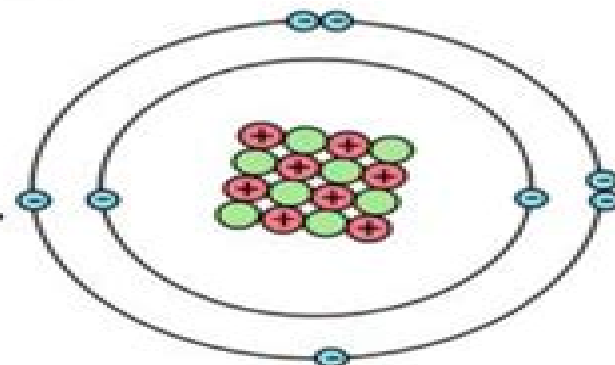


## **Atomic Theory – Unit Exam**

**For all of the following multiple-choice questions, choose only one correct answer:**

1. Which description below accurately describes the current model of the atom?
  - a) Electrons are distributed around the nucleus and have almost no mass compared to protons
  - b) Protons, electrons and neutrons are evenly distributed throughout all parts of the atom
  - c) The nucleus consists of protons, electrons and neutrons
  - d) The nucleus is made of electrons and protons
2. Which of the following is **true** about subatomic particles?
  - a) Electrons are negatively charged and are the heaviest subatomic particle
  - b) The mass of a proton nearly equals the mass of a neutron
  - c) Protons are positively charged and the lightest subatomic particle
  - d) Neutrons are negatively charged and are the heaviest subatomic particle
3. What would be the **mass** of the atom pictured to the right?
  - a) 7 amu
  - b) 8 amu
  - c) 16 amu
  - d) 24 amu
4. How many **valence electrons** are shown in the atom on the right?
  - a) 2 valence electrons
  - b) 4 valence electrons
  - c) 6 valence electrons
  - d) 8 valence electrons
5. Why are atoms electrically **neutral**?
  - a) The number of neutrons are equal to the number of protons and electrons.
  - b) The numbers of neutrons and electrons are equal and cancel each other out.
  - c) The numbers of protons and electrons are equal and cancel each other out.
  - d) The numbers of protons and neutrons are equal and cancel each other out.
6. Based on the symbol to the right, what can be concluded about the element **fluorine**?
  - a) There are 19 atoms in a molecule of fluorine
  - b) There are 9 neutrons outside of the fluorine atom's nucleus
  - c) There are 9 protons in the nucleus of a fluorine atom
  - d) There are 9 valence electrons in the fluorine atom.



# Electron Structure Test Review Guide

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