

3. Refer to the data table in Question 2.

a. How do the reduced ratios in the last column compare to the coefficients in the reaction shown in Model 1?

*The ratios are the same as the coefficients in Model 1: 1:3:2.*

b. Use mathematical concepts to explain how your answer in part a is possible.

*The coefficients were all multiplied by the same number, so they all ended back to the same thing.*

4. Even 538 is a small number of molecules to use in a reaction. Typically chemists use much larger numbers of molecules. (Recall that one mole is equal to  $6.02 \times 10^{23}$  particles.) Consider each situation below as it relates to the reaction in Model 1:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ .

a. Calculate the amount of reactants consumed and products made.

b. Record the ratio of  $\text{N}_2$  to  $\text{H}_2$  to  $\text{NH}_3$ . Reduce the ratio to the lowest whole number possible.

	$\text{N}_2$ Consumed	$\text{H}_2$ Consumed	$\text{NH}_3$ Produced	Ratio $\text{N}_2:\text{H}_2:\text{NH}_3$
If the reaction occurred $6.02 \times 10^{23}$ times, how many molecules would be consumed or produced?	$6.02 \times 10^{23}$	$1.81 \times 10^{24}$	$1.20 \times 10^{24}$	1 : 3 : 2
How many <i>moles</i> of each substance would be consumed or produced in the previous situation?	1	3	2	1 : 3 : 2

5. Refer to the data table in Question 4.

a. How do the reduced ratios in the last column compare to the coefficients in the reaction in Model 1?

*The reduced ratios are the same as the coefficients.*

b. Use mathematical concepts to explain how your answer in part a is possible.

*We are still multiplying coefficients by the same number so we just didn't everything by the number to get back to a reduced ratio.*

6. The ratio obtained from the coefficients in a balanced chemical equation is called the **mole ratio**.

a. What is the mole ratio for the reaction in Model 1?



b. Explain why this ratio is called the mole ratio?

*It is the reduced ratio in moles (all the coefficients =  $6.02 \times 10^{23}$ ).*

# Mole Ratios Pogil Activities For High School Chemistry

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