

CHAPTER 19 REVIEW

Oxidation-Reduction Reactions**SECTION 3****SHORT ANSWER** Answer the following questions in the space provided.

1. For each of the following, identify the stronger oxidizing or reducing agent. (Refer to **Figure 3.2** of the text.)

_____ a. Ca or Cu as a reducing agent
_____ b. Ag^+ or Na^+ as an oxidizing agent
_____ c. Fe^{3+} or Fe^{2+} as an oxidizing agent
2. For each of the following incomplete equations, state whether a redox reaction is likely to occur. (Refer to **Figure 3.2** of the text.)

_____ a. $\text{Mg} + \text{Sn}^{2+} \rightarrow$
_____ b. $\text{Ag} + \text{Cu}^{2+} \rightarrow$
_____ c. $\text{Br}_2 + \text{I}^- \rightarrow$
3. Label each of the following statements about redox as True or False.

_____ a. A strong oxidizing agent is itself readily reduced.
_____ b. In disproportionation, one chemical acts as both an oxidizing agent and a reducing agent in the same process.
_____ c. The number of moles of chemical oxidized must equal the number of moles of chemical reduced.
4. Solutions of Fe^{2+} are fairly unstable, in part because they can undergo disproportionation, as shown by the following unbalanced equation:
$$\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{Fe}$$
 - a. Balance the above equation.

 - b. If the reaction described above produces 0.036 mol of Fe, how many moles of Fe^{3+} will form?

Modern Chemistry Oxidation Reduction Reactions

Section Review

Clifford Lane



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