

CHAPTER 17

RADIOACTIVITY AND NUCLEAR CHEMISTRY

Practice Problems

1. (a) Write a nuclear equation for alpha decay of Po-210 .



- (b) Write a nuclear equation for the beta decay of Th-232 .



- (c) Write a nuclear equation for positron emission by Co-57 .



2. A 10.0-g sample of a radioactive nuclide is administered to a patient for cancer diagnosis. If the nuclide has a half-life of 1.5 hours, how much of the nuclide remains in the patient after 4.5 days?

$$\text{4.5 days} \times \frac{24 \text{ hours}}{1 \text{ day}} \times \frac{60 \text{ min}}{1 \text{ hour}} = 6480 \text{ min} \quad \frac{6480 \text{ min}}{90 \text{ min}} = 72 \text{ half-lives}$$

- (a) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (b) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (c) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (d) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (e) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (f) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (g) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (h) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (i) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$ (j) $10.0 \text{ g} \times \left(\frac{1}{2}\right)^{72} = 1.5 \times 10^{-21} \text{ g}$

3. A mammoth skeleton has a carbon-14 content of 1.56% of that found in living organisms. When did the mammoth live?

The age of the mammoth is 14,400 years. The half-life of C-14 is 5,730 years.

$$\frac{1.56\%}{100\%} = \left(\frac{1}{2}\right)^n \quad n = 4 \quad 4 \times 5,730 \text{ years} = 22,920 \text{ years}$$

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RC Schank



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Table of Contents Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers

1. Understanding the eBook Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - The Rise of Digital Reading Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Personalized Recommendations
 - Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers User Reviews and Ratings
 - Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers and Bestseller Lists
5. Accessing Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers Free and Paid eBooks
 - Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers Public Domain eBooks
 - Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers eBook Subscription Services
 - Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers Budget-Friendly Options

6. Navigating Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers Compatibility with Devices
 - Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Highlighting and Note-Taking Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Interactive Elements Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
8. Staying Engaged with Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
9. Balancing eBooks and Physical Books Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Setting Reading Goals Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Fact-Checking eBook Content of Modern Chemistry Chapter 21 Review Nuclear Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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