

$$[\text{H}^+] = 0.01$$

$$\text{pH} = -\log_{10} [\text{H}^+]$$

$$\text{pH} = -\log_{10} [0.01]$$

$$\text{pH} = -\log_{10} [1 \times 10^{-2}]$$

$$\text{pH} = -(\log_{10} 1 - 2 \log_{10} 10)$$

$$\text{pH} = 0 + 2 \times \log_{10} 10$$

$$\text{pH} = 0 + 2 \times 1 = 2$$

$$\text{pH} = 2$$

SamacheerKalvi.Guru

Modern Chemistry Chapter 18 Review Chemical Equilibrium Answers

Michael Hoelscher



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