

### Limiting Reactant Practice 2.3 - Answer Key

**Objective:** Use stoichiometry to identify the limiting and excess reactants.

**Directions:** Answer the following limiting reactant problems. Show your work!!

1. A reaction between potassium and oxygen gas produces potassium oxide, represented by the following balanced chemical equation:



- a. If 3.00 g of potassium and 5.00 g of oxygen are mixed, how much potassium oxide is produced?

$$3.00 \text{ g K} \times \left( \frac{1 \text{ mol K}}{39.10 \text{ g K}} \right) \left( \frac{2 \text{ mol K}_2\text{O}}{4 \text{ mol K}} \right) \left( \frac{94.2 \text{ g K}_2\text{O}}{1 \text{ mol K}_2\text{O}} \right) = 11.1 \text{ g K}_2\text{O}$$

$$5.00 \text{ g O}_2 \times \left( \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right) \left( \frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) \left( \frac{94.2 \text{ g K}_2\text{O}}{1 \text{ mol K}_2\text{O}} \right) = 29.4 \text{ g K}_2\text{O}$$

- b. What is the limiting reactant?

Potassium (K)

- c. What is the reactant in excess?

Oxygen (O<sub>2</sub>)

2. In a reaction, 12.00 g of aqueous sodium hydroxide reacts with aqueous 25.00 g of copper(II) sulfate.



- a. What mass of sodium sulfate is produced?

$$12.00 \text{ g NaOH} \times \left( \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \right) \left( \frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol NaOH}} \right) \left( \frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 17.10 \text{ g}$$

$$25.00 \text{ g CuSO}_4 \times \left( \frac{1 \text{ mol CuSO}_4}{159.61 \text{ g CuSO}_4} \right) \left( \frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) \left( \frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 28.63 \text{ g}$$

- b. What is the limiting reactant?

Copper(II) sulfate (CuSO<sub>4</sub>)

- c. What is the reactant in excess?

Sodium Hydroxide (NaOH)

# Limiting Reactants Study Guide Answers

**John A. Olmsted, Gregory M. Williams**



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**CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE** Kevin Smith, 2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text

*Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

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[Organic Chemistry Study Guide with Solutions Manual](#) Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

**General Chemistry Workbook** Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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