

Limiting Reactant Practice 2.3 - Answer Key

Objective: Use stoichiometry to identify the limiting and excess reactants.

Directions: Answer the following limiting reactant problems. Show your work!!

1. A reaction between potassium and oxygen gas produces potassium oxide, represented by the following balanced chemical equation:



- a. If 3.00 g of potassium and 5.00 g of oxygen are mixed, how much potassium oxide is produced?

$$3.00 \text{ g K} \times \left(\frac{1 \text{ mol K}}{39.10 \text{ g K}} \right) \left(\frac{1 \text{ mol O}_2}{2 \text{ mol K}} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = \boxed{1.21 \text{ g K}_2\text{O}}$$

$$5.00 \text{ g O}_2 \times \left(\frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right) \left(\frac{1 \text{ mol K}}{2 \text{ mol O}_2} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = 1.56 \text{ g K}_2\text{O}$$

- b. What is the limiting reactant?

Potassium (K)

- c. What is the reactant in excess?

Oxygen (O₂)

2. In a reaction, 12.00 g of aqueous sodium hydroxide reacts with aqueous 22.00 g of copper(II) sulfate.



- a. What mass of sodium sulfate is produced?

$$12.00 \text{ g NaOH} \times \left(\frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol NaOH}} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 17.10 \text{ g}$$

$$22.00 \text{ g CuSO}_4 \times \left(\frac{1 \text{ mol CuSO}_4}{159.61 \text{ g CuSO}_4} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 19.67 \text{ g}$$

- b. What is the limiting reactant?

Copper(II) sulfate (CuSO₄)

- c. What is the reactant in excess?

Sodium Hydroxide (NaOH)

Limiting Reactants Study Guide Answer Key

John A. Olmsted, Gregory M. Williams



Limiting Reactants Study Guide Answer Key:

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