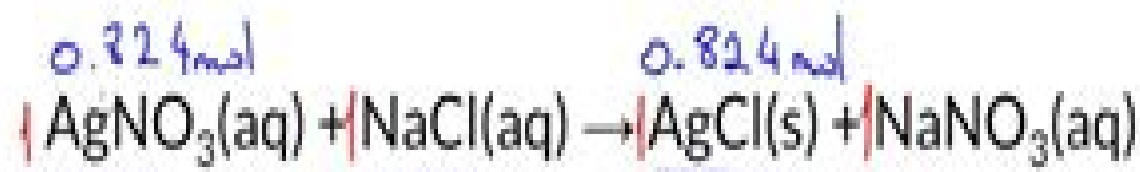


When 140.0 g of  $\text{AgNO}_3$  was added to an aqueous solution of  $\text{NaCl}$ , 86.0 g of  $\text{AgCl}$  was collected as a white precipitate. Which salt was the limiting reactant in this reaction? How many grams of  $\text{NaCl}$  was present in the solution?



1) L.R. - ?

2)  $m(\text{NaCl})$  - ?

$$n(\text{AgNO}_3) = \frac{m}{M} = \frac{140 \text{ g}}{169.9 \text{ g/mol}} = 0.824 \text{ mol}$$

$$n(\text{AgCl}) = \frac{m}{M} = \frac{86.0 \text{ g}}{143.3 \text{ g/mol}} = \underline{0.600 \text{ mol A.Y.}}$$

# Limiting Reactant Practice Problems With Answers

**Dr. B. Kapila, S. K. Khanna**



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