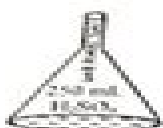


অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার

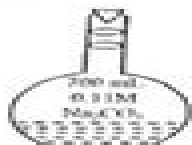


পরীক্ষার কপুড়া, এখানে কোর্স পরীক্ষা, আরও কয়েকজন শিক্ষার্থীর কয়েকজনমুকের পরীক্ষা এবং মতামত দেওয়ার পরামর্শ লুপাৎ উক্ত অধ্যায়ভিত্তিক মেঞ্জর হয়েছে। এগুলোর অনুশীলন করলে তুমি এ অধ্যায়টি থেকে পেয়েছো সৃজনশীল জ্ঞানমূলক প্রশ্নের উত্তর লিপিতে পারবে সহজেই।

প্রশ্ন



পাত্র-A



পাত্র-B

সি. কে. ১০/১৪/১৭

- ক. মোলার প্রবণ কী?
 খ. "কোনস পানিদ্রুত লবণের দ্রাব্যতা, তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।
 গ. A-পাত্রে বিদ্যমান এলিডের ভর কত নির্ণয় করো।
 ঘ. ট্রবল দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা পরিষ্কৃত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

১০০০ mL দ্রবণে কোন দ্রবের 1 mol দ্রবীভূত থাকলে উক্ত দ্রবকে মোলার দ্রবণ বলে।

কোনস পানিদ্রুত লবণের দ্রাব্যতা তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে দ্রাব্যতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কোলাস পানি দ্রুত লবণ নিষ্কৃতি হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে দ্রবীভূত করার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরূপী বিক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে দ্রাব্যতার হ্রাস ঘটে।

১

আমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
 অতএব, দ্রবীভূত H_2SO_4 এর ভর = 12.25 g

উদ্দীপকের A-পাত্রে অবস্থিত, H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
 দ্রবণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
 দ্রবণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$
 দ্রবীভূত H_2SO_4 এর ভর, $w = ?$

১০) পাত্র-A: H_2SO_4 এর মোল সংখ্যা = $\left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

পাত্র-B: Na_2CO_3 এর মোল সংখ্যা = $\left(0.11 \times \frac{200}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্ভবতঃ বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$
 সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$
 1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে 0.077 mol "

অনশিষ্ট H_2SO_4 এর পরিমাণ = $(0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লীয়।

১১) একমল ঘনত্ব ল্যাবরেটরিতে X-পদ্ধতিতে 10 mL এর ত্রেসে প পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর মূল্য নির্ণয় করো।

- ক. নোট কাকে বলে?
 খ. Cl_2 অপোলার— ব্যাখ্যা করো।
 গ. উদ্দীপকের এলিডটির মোলার ঘনমাত্রা নির্ণয় করো।
 ঘ. উদ্দীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

১২) নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

১৩) সমযোজী বন্ধনে আকর্ষণ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা কম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহূণ ঘ থাকে।

১৪) উদ্দীপকে বর্ণিত এলিডটি হলো 2% $\left(\frac{W}{V}\right) \text{ HCl}$ ।

2% $\left(\frac{W}{V}\right) \text{ HCl}$ এর অর্থ হলো—
 100 mL দ্রবণে HCl এর ভর 2g

∴ 1000 mL " " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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