

(ix) The standard potential of the cell in the following reaction is _____



($E_{\text{Cd}^{2+}}^{\circ} = -0.403 \text{ V}$, $E_{\text{Cu}^{2+}}^{\circ} = 0.334 \text{ V}$)

(a) -0.737 V

(b) 0.737 V

(c) -0.069 V

(d) 0.069 V

(x) The value of $[\text{H}_3\text{O}^+]$ in mol lit^{-1} of 0.001 M acetic acid solution ($K_a = 1.8 \times 10^{-5}$) is _____

(a) 1.34×10^{-1}

(b) 1.34×10^{-2}

(c) 1.34×10^{-3}

(d) 1.34×10^{-4}

Q.2. Answer the following questions:

[8]

(i) Write the product formed when alkyl halide reacts with silver nitrite.

(ii) Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.

(iii) Write the name of biodegradable polyamide copolymer.

(iv) Identify the molecularity of following elementary reaction:



(v) What is the action of selenium on magnesium metal?

(vi) Write the name of isomerism in the following complexes:



(vii) Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.

(viii) Henry's law constant for $\text{CH}_3\text{Br}_{(\text{g})}$ is $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$ at 25°C . What is solubility of $\text{CH}_3\text{Br}_{(\text{g})}$ in water at same temperature and partial pressure of 0.164 bar ?

SECTION – B

Attempt any EIGHT of the following questions:

[16]

Q.3. Explain pseudo-first order reaction with suitable example.

Q.4. Write the consequences of Schottky defect with reasons.

Q.5. What is the action of following on ethyl bromide:

(i) Na in dry ether

(ii) Mg in dry ether

Q.6. Explain formation of peptide linkage in protein with an example.

Q.7. Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.

Q.8. Explain monodentate and ambidentate ligands with example.

Q.9. Explain the trends in the following atomic properties of group 16 elements:

(i) Atomic radii

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Mike Jess



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Table of Contents Hsc Board Question Papers Science

1. Understanding the eBook Hsc Board Question Papers Science
 - The Rise of Digital Reading Hsc Board Question Papers Science
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Board Question Papers Science
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Board Question Papers Science
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Board Question Papers Science
 - Personalized Recommendations
 - Hsc Board Question Papers Science User Reviews and Ratings
 - Hsc Board Question Papers Science and Bestseller Lists
5. Accessing Hsc Board Question Papers Science Free and Paid eBooks
 - Hsc Board Question Papers Science Public Domain eBooks
 - Hsc Board Question Papers Science eBook Subscription Services
 - Hsc Board Question Papers Science Budget-Friendly Options
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 - Hsc Board Question Papers Science Compatibility with Devices
 - Hsc Board Question Papers Science Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Board Question Papers Science
 - Highlighting and Note-Taking Hsc Board Question Papers Science
 - Interactive Elements Hsc Board Question Papers Science
8. Staying Engaged with Hsc Board Question Papers Science

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Hsc Board Question Papers Science
- 9. Balancing eBooks and Physical Books Hsc Board Question Papers Science
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Board Question Papers Science
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Board Question Papers Science
 - Setting Reading Goals Hsc Board Question Papers Science
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Board Question Papers Science
 - Fact-Checking eBook Content of Hsc Board Question Papers Science
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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