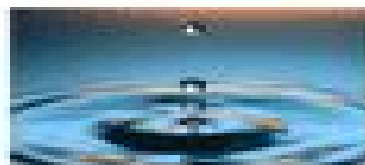


Name: _____

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Intermolecular Forces Lab Worksheet - Answer Key



Directions: Today you will complete two experiments that compare the properties of water to isopropanol (a type of rubbing alcohol). Before conducting each experiment, you will first read about the experiment and then make a prediction about what you think will happen. Afterward you do the experiment, you will write an explanation of what was happening at the atomic level that let us observe these properties.

Introduction: In many ways, water is a miracle liquid. It is essential for all living things on earth and it is often referred to as a universal solvent because many substances dissolve in it. Water displays unusual properties due to the ways the individual water molecules interact with each other.

1. Draw the structural formula (with lone pairs) for the two molecules we will be comparing today, water (H_2O) and isopropanol ($\text{C}_3\text{H}_7\text{O}$). Draw a dipole arrow to show the direction the electrons are being pulled and label the areas of partial positive and negative charges.

H_2O



or



$\text{C}_3\text{H}_7\text{O}$

Note: the bond between C and H is NOT POLAR. Do not draw dipole arrows or partial charges for those bonds.



2. Which molecule is more polar? water

Experiment #1: Penny

How many drops of water fit on the head of a penny? How many drops of isopropanol fit on the head of a penny?

Prediction:

- a. I think _____ drops of water will fit on the head of the penny because _____

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G Thomas



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Table of Contents Grade 11 Intermolecular Forces Experiment Solutions

1. Understanding the eBook Grade 11 Intermolecular Forces Experiment Solutions
 - The Rise of Digital Reading Grade 11 Intermolecular Forces Experiment Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 11 Intermolecular Forces Experiment Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 11 Intermolecular Forces Experiment Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 11 Intermolecular Forces Experiment Solutions
 - Personalized Recommendations
 - Grade 11 Intermolecular Forces Experiment Solutions User Reviews and Ratings

- Grade 11 Intermolecular Forces Experiment Solutions and Bestseller Lists
- 5. Accessing Grade 11 Intermolecular Forces Experiment Solutions Free and Paid eBooks
 - Grade 11 Intermolecular Forces Experiment Solutions Public Domain eBooks
 - Grade 11 Intermolecular Forces Experiment Solutions eBook Subscription Services
 - Grade 11 Intermolecular Forces Experiment Solutions Budget-Friendly Options
- 6. Navigating Grade 11 Intermolecular Forces Experiment Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Grade 11 Intermolecular Forces Experiment Solutions Compatibility with Devices
 - Grade 11 Intermolecular Forces Experiment Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 11 Intermolecular Forces Experiment Solutions
 - Highlighting and Note-Taking Grade 11 Intermolecular Forces Experiment Solutions
 - Interactive Elements Grade 11 Intermolecular Forces Experiment Solutions
- 8. Staying Engaged with Grade 11 Intermolecular Forces Experiment Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 11 Intermolecular Forces Experiment Solutions
- 9. Balancing eBooks and Physical Books Grade 11 Intermolecular Forces Experiment Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 11 Intermolecular Forces Experiment Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grade 11 Intermolecular Forces Experiment Solutions
 - Setting Reading Goals Grade 11 Intermolecular Forces Experiment Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade 11 Intermolecular Forces Experiment Solutions
 - Fact-Checking eBook Content of Grade 11 Intermolecular Forces Experiment Solutions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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