

Molarity Practice problems

$$M = \frac{\text{mol}}{L}$$



Key

1. How many moles of CaCl_2 would be dissolved in 1.0 L of a 0.1 M solution of CaCl_2 ?

$$\text{mol} = M \times L = 0.1 \frac{\text{mol}}{L} \times 1 L = 0.1 \text{ mol } \text{CaCl}_2$$

2. What is the concentration of 25 mL of solution that has 0.05 moles of solute dissolved in it?

$$M = \frac{\text{mol}}{L} = \frac{0.05 \text{ mol}}{0.025 L} = 2 M$$

3. How many liters of water need to be added to 0.5 moles of CaCl_2 to make a 0.20 M solution of CaCl_2 ?

$$L = \frac{\text{mol}}{M} = \frac{0.5 \text{ mol}}{0.20 M} = 2.5 L$$

4. How many grams of potassium carbonate are needed to make 200 mL of a 2.5 M solution?

$$\text{mol} = M \times L \quad \text{K}_2\text{CO}_3$$

$$\text{mol} = 2.5 M \times 0.200 L = 0.50 \text{ mol}$$

$$\frac{0.50 \text{ mol } \text{K}_2\text{CO}_3}{1 \text{ mol}} \times 138.21 \text{ g} = 69.1 \text{ g } \text{K}_2\text{CO}_3$$

5. How many grams of Lithium Bromide would be needed to make 1.43 L of 4 M Lithium Bromide solution?

$$\text{mol} = M \times L \quad \text{LiBr}$$

$$= 4 M \times 1.43 L = 5.72 \text{ mol } \text{LiBr}$$

$$\frac{5.72 \text{ mol } \text{LiBr}}{1 \text{ mol}} \times 86.84 \text{ g} = 496.7 \text{ g } \text{LiBr}$$

6. What is the concentration of a 450 mL solution that contains 200 grams of iron (II) chloride? FeCl_2

$$M = \frac{\text{mol}}{L}$$

$$\frac{200 \text{ g } \text{FeCl}_2}{126.75 \text{ g}} \times 1 \text{ mol} = 1.578 \text{ mol } \text{FeCl}_2$$

$$\frac{1.578 \text{ mol}}{0.450 L} = 3.5 M$$

7. How many grams of ammonium sulfate are needed to make a 0.25 L solution at a concentration of 6 M?

$$\text{mol} = M \times L \quad (\text{NH}_4)_2\text{SO}_4$$

$$= 6 M \times 0.25 L = 1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4$$

$$\frac{1.5 \text{ mol } (\text{NH}_4)_2\text{SO}_4}{1 \text{ mol}} \times 132.138 \text{ g} = 198.2 \text{ g } (\text{NH}_4)_2\text{SO}_4$$

8. What is the concentration of a solution that has a volume of 2.5 L and contains 660 grams of calcium phosphate?

$$M = \frac{\text{mol}}{L}$$

$$\frac{660 \text{ g } \text{Ca}_3(\text{PO}_4)_2}{310.18 \text{ g}} \times 1 \text{ mol} = 2.127 \text{ mol}$$

$$\frac{2.127 \text{ mol}}{2.5 L} = 0.85 M$$

9. How many grams of copper (II) fluoride are needed to make 6.7 liters of a 1.2 M solution?

$$\text{mol} = M \times L \quad \text{CuF}_2$$

$$= 1.2 M \times 6.7 L = 8.04 \text{ mol } \text{CuF}_2$$

$$\frac{8.04 \text{ mol } \text{CuF}_2}{1 \text{ mol}} \times 101.55 \text{ g} = 816.5 \text{ g } \text{CuF}_2$$

Molarity Practice Unit 10 Lesson Key

Lingjun Ying



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Table of Contents Molarity Practice Unit 10 Lesson Key

1. Understanding the eBook Molarity Practice Unit 10 Lesson Key
 - The Rise of Digital Reading Molarity Practice Unit 10 Lesson Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Practice Unit 10 Lesson Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Practice Unit 10 Lesson Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Practice Unit 10 Lesson Key
 - Personalized Recommendations
 - Molarity Practice Unit 10 Lesson Key User Reviews and Ratings
 - Molarity Practice Unit 10 Lesson Key and Bestseller Lists

5. Accessing Molarity Practice Unit 10 Lesson Key Free and Paid eBooks
 - Molarity Practice Unit 10 Lesson Key Public Domain eBooks
 - Molarity Practice Unit 10 Lesson Key eBook Subscription Services
 - Molarity Practice Unit 10 Lesson Key Budget-Friendly Options
6. Navigating Molarity Practice Unit 10 Lesson Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity Practice Unit 10 Lesson Key Compatibility with Devices
 - Molarity Practice Unit 10 Lesson Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity Practice Unit 10 Lesson Key
 - Highlighting and Note-Taking Molarity Practice Unit 10 Lesson Key
 - Interactive Elements Molarity Practice Unit 10 Lesson Key
8. Staying Engaged with Molarity Practice Unit 10 Lesson Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Practice Unit 10 Lesson Key
9. Balancing eBooks and Physical Books Molarity Practice Unit 10 Lesson Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Practice Unit 10 Lesson Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molarity Practice Unit 10 Lesson Key
 - Setting Reading Goals Molarity Practice Unit 10 Lesson Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Practice Unit 10 Lesson Key
 - Fact-Checking eBook Content of Molarity Practice Unit 10 Lesson Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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