



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p>Equilibrium Constant</p> <p>The equilibrium constant, K, is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	
<p>Equilibrium Constant, K_c</p> <p>K_c is the equilibrium constant based on concentrations. It is used for reactions in solution or in the gas phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_p</p> <p>K_p is the equilibrium constant based on partial pressures. It is used for reactions in the gas phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	
<p>Equilibrium Constant, K</p> <p>K is the equilibrium constant based on activities. It is used for reactions in any phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{a_C^c a_D^d}{a_A^a a_B^b}$	<p>Relationship between K_c and K_p</p> <p>The relationship between K_c and K_p is given by the equation:</p> $K_p = K_c (RT)^{\Delta n}$ <p>where R is the gas constant, T is the temperature in Kelvin, and Δn is the change in the number of moles of gas.</p>	
<p>Le Chatelier's Principle: the effect of changing conditions on equilibrium</p>		
<p>Concentration</p>	<p>Increasing the concentration of a reactant or product shifts the equilibrium in the direction that consumes that species.</p>	<p>For the reaction $N_2 + 3H_2 \rightleftharpoons 2NH_3$, increasing the concentration of N_2 shifts the equilibrium to the right, producing more NH_3.</p>
<p>Pressure</p>	<p>Increasing the pressure shifts the equilibrium in the direction that produces fewer moles of gas.</p>	<p>For the reaction $N_2 + 3H_2 \rightleftharpoons 2NH_3$, increasing the pressure shifts the equilibrium to the right, producing more NH_3.</p>
<p>Temperature</p>	<p>Increasing the temperature shifts the equilibrium in the direction that absorbs heat (endothermic reaction).</p>	<p>For the reaction $N_2 + 3H_2 \rightleftharpoons 2NH_3$, increasing the temperature shifts the equilibrium to the left, producing less NH_3.</p>

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Julia Schneider



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