# **Chemistry Chapter 3**

# Types of bonds and Naming compounds

#### **Chemical Bonds**

- Compounds are made of atoms held together by chemical bonds
- Ionic bonds: Electrostatic interaction formed by oppositely charged ions
- Covalent bonds result when two atoms share some of their electrons

# Naming simple compounds (a)

- Rules for Binary Ionic Compounds:
  - Cation first, then anion (CA)
  - Cation is named after the element
  - Anion name by taking root of chemical name and adding side to it
- Examples
  - o CaS
  - o KF
  - o MgO

#### Naming simple compounds (b)

- Rules for Binary Ionic Compounds (Transition metals):
  - Cation first, then anion (C()A)
  - Cation is named after the element, but include charge
  - Anion name by taking root of chemical name and adding ide to it
- Examples
  - FeCl<sub>3</sub> Iron (III) chloride
  - FeCl<sub>2</sub> Iron (II) chloride
  - Cu<sub>2</sub>O Copper (I) Oxide
  - CuO Copper (II) Oxide

# Naming simple compounds (c)

- Rules for Molecular Compounds (Type C):
  - The first element named fully
  - Second named as if it were an anion
  - Prefixes are denoted to specify number of atoms, except when there is just one of the first element
- Examples
  - H<sub>2</sub>O Dihydrogen monoxide
  - CO Carbon Monoxide
  - CO<sub>2</sub> Carbon Dioxide

# **Modern Chemistry Chapter 3 Notes**

**B** Lingard

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