

Recitation Worksheet Six

Name:

Key

UGA ID:

Instructions:

- Please enter your first and last name as it appears on the eLC roster (do not use a nickname that is not reflected in eLC).
- Your UGA myID is a combination of letters and numbers (example: mine is jmj81738). Do not enter your 81x number.
- Download this worksheet and print it if you have a printer. Write the answers in the answer boxes and show your work when appropriate. Using the instructions in the Welcome module on eLC, convert your worksheet to a PDF and then upload it to Gradescope. If you have an iPhone or Android device, you can scan and upload directly through the Gradescope app. The pages must be in the correct order or Gradescope will not be able to read it.
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- Answers must be written in the corresponding answer box or no credit will be awarded.
- This worksheet is due no later than 11:59 PM on the Friday of the recitation week.
- The instructions for uploading worksheets to Gradescope can be found in the Content area of eLC in the Welcome Module.
- You must show your work to receive credit.

1. If 0.50 moles of nitrogen gas reacts with 0.50 moles of oxygen gas, how many molecules of nitrogen dioxide are produced?

3.0E23

molecules



$$0.50 \text{ mol } N_2 \times \frac{2 \text{ mol } NO_2}{1 \text{ mol } N_2} \times \frac{6.022 \times 10^{23} \text{ molecules}}{1 \text{ mol } NO_2} = 6.022 \times 10^{23} \text{ molecules } NO_2$$

limiting reagent

$$0.50 \text{ mol } O_2 \times \frac{2 \text{ mol } NO_2}{2 \text{ mol } O_2} \times \frac{6.022 \times 10^{23} \text{ molecules}}{1 \text{ mol}} = 3.011 \times 10^{23} \text{ molecules } NO_2$$

Limiting Reagent Section Review Key

Michal Rosen-Zvi



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