

Limiting Reactant

For the following reactions, find the following:

- Which of the reactants is the limiting reactant?
 - What is the maximum amount of each product that can be formed?
 - How much of the excess reactant is left over after the reaction is complete?
- See next page for calculations!

1. Consider the following reaction:



a) limiting reactant

Answer the questions above, assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate.

$$\frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \left| \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \right| \left| \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{3 \text{ mol NH}_4\text{NO}_3} \right| \left| \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} \right| = 18.6 \text{ g } (\text{NH}_4)_3\text{PO}_4$$

$$\frac{50.0 \text{ g Na}_3\text{PO}_4}{1} \left| \frac{1 \text{ mol Na}_3\text{PO}_4}{163.94 \text{ g Na}_3\text{PO}_4} \right| \left| \frac{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol Na}_3\text{PO}_4} \right| \left| \frac{149.12 \text{ g } (\text{NH}_4)_3\text{PO}_4}{1 \text{ mol } (\text{NH}_4)_3\text{PO}_4} \right| = 45.5 \text{ g } (\text{NH}_4)_3\text{PO}_4$$

$$\frac{30.0 \text{ g NH}_4\text{NO}_3}{1} \left| \frac{1 \text{ mol NH}_4\text{NO}_3}{80.0 \text{ g NH}_4\text{NO}_3} \right| \left| \frac{3 \text{ mol NaNO}_3}{3 \text{ mol NH}_4\text{NO}_3} \right| \left| \frac{85.00 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} \right| = 31.9 \text{ g NaNO}_3$$

2. Consider the following reaction:



a) limiting reactant

Answer the questions at the top of this sheet, assuming we start with 100.0 grams of calcium carbonate and 45.0 grams of iron (II) phosphate.

$$\frac{100.0 \text{ g CaCO}_3}{1} \left| \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \right| \left| \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCO}_3} \right| \left| \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} \right| = 103.3 \text{ g Ca}_3(\text{PO}_4)_2$$

$$\frac{45.0 \text{ g FePO}_4}{1} \left| \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \right| \left| \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol FePO}_4} \right| \left| \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} \right| = 46.3 \text{ g Ca}_3(\text{PO}_4)_2$$

$$\frac{45.0 \text{ g FePO}_4}{1} \left| \frac{1 \text{ mol FePO}_4}{150.82 \text{ g FePO}_4} \right| \left| \frac{1 \text{ mol Fe}_2(\text{CO}_3)_3}{2 \text{ mol FePO}_4} \right| \left| \frac{291.73 \text{ g Fe}_2(\text{CO}_3)_3}{1 \text{ mol Fe}_2(\text{CO}_3)_3} \right| = 43.5 \text{ g Fe}_2(\text{CO}_3)_3$$

Limiting Reactant Packet

Puja Mehta



Limiting Reactant Packet:

Limiting Reactant Packet Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Limiting Reactant Packet**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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