

LAB #5 - LIMITING REAGENT**Objective**

- Use stoichiometry to determine the limiting reactant.
- Calculate the theoretical yield.
- Calculate the percent yield of a reaction.

Introduction

In lecture you have learned to read chemical equations and evaluate the mol to mol ratios of reactants and products involved in a chemical reaction. In laboratory experiments it is difficult to measure out chemicals in the exact ratio necessary for the chemical reaction. For time and speed reasons, the reaction mixtures in lab will usually have a limiting and an excess reactant.

Limiting reagent (also called limiting reactant) problems use stoichiometry to determine the theoretical yield for a chemical reaction. The *limiting reactant* will be completely consumed in the reaction and limits the amount of product you can make. The limiting reactant also determines the amount of product you can make (the *theoretical yield*). The reactant that is left over after the reaction is complete is called the *excess reactant*.

Example 1:

Consider the process it takes to make a ham sandwich. You need two slices of bread and one piece of ham to make each sandwich. How many complete sandwiches could you make if you had eighteen slices of bread and six slices of ham? Let's set this up like a chemical equation where the ham and bread are our reactants and the sandwich is our product:



Now we can use stoichiometry to determine the amount of sandwiches we could make if we used all of our reactants.

$$18 \text{ slices of bread} \times \frac{1 \text{ sandwich}}{2 \text{ pieces of bread}} = 9 \text{ sandwiches}$$

$$6 \text{ slices of ham} \times \frac{1 \text{ sandwich}}{1 \text{ slice of ham}} = 6 \text{ sandwiches}$$

Limited Reagent Lab Key

David A. Ucko

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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