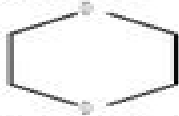


F325 Past paper Answers

1. (a) G: CO (1)
 $\text{HCOOH}/\text{H}_2\text{SO}_4 \rightarrow \text{CO} + \text{H}_2\text{O}$ (1) 2
- (b) H: C (1)
 $\text{C}_{12}\text{H}_{22}\text{O}_{11} \rightarrow 12\text{C} + 11\text{H}_2\text{O}$ (1) 2
- (c) I: $\text{C}_4\text{H}_8\text{O}_2$ (1)
 $2\text{C}_2\text{H}_4\text{O}_2 \rightarrow \text{C}_4\text{H}_8\text{O}_2 + 2\text{H}_2\text{O}$ (1)
 Structure:
 (1)
 accept any sensible structure of $\text{C}_4\text{H}_8\text{O}_2$ 3
2. (a) empirical formula
 $\text{N} : \text{O} = 63.64/14 : 36.36/16$ (1)
 $= 4.56 : 2.27 = 2 : 1$. Empirical formula = N_2O (1)
 molecular formula
 M_r of gas = $1.833 \times 24 = 44$ (1)(calc 43.992)
 with these two pieces of evidence, assume that molecular formula = N_2O 3
- (b) any chemical that reacts to produce gas:
 e.g. carbonate and CO_2 (1)
 accept: metal more reactive than Pb and H_2
 balanced equation to match chemical added (1) 2
- (c) $M_r(\text{Lidocaine}) = 236$ (1)
 moles Novocaine = $100 \times 10^{-3}/236 = 4.24 \times 10^{-4}$ (1)
 concentration of Novocaine = $4.24 \times 10^{-4} \times (1000/5)$
 $= 0.0847/0.0848 \text{ mol dm}^{-3}$ (1) 3
- (d) mass C = $12 \times 3.74/44.0 = 12 \times 0.085 = 1.02 \text{ g}$ (1)
 mass H = $2/18 \times 0.918 = 0.102 \text{ g}$ (1)
 mass O = $1.394 - (1.020 + 0.102) = 0.272 \text{ g}$
 ratio C : H : O = $1.02/12 : 0.102/1 : 0.272/16$ (1)
 $= 0.0850 : 0.102 : 0.0170 / 5 : 6 : 1 / \text{C}_5\text{H}_6\text{O}$ (1)
 $\text{C}_5\text{H}_{10}\text{O}$ has relative mass of 82
 M_r is 164 = 2×82
 \therefore molecular formula = $\text{C}_{10}\text{H}_{12}\text{O}_2$ (1) 5
3. (i) $4\text{PCl}_3 + 10\text{MgO} \rightarrow \text{P}_4\text{O}_{10} + 10\text{MgCl}_2$ (1) 1
- (ii) $100\text{g P}_4\text{O}_{10} = \frac{100}{284} / 0.35(2) \text{ mol}$ (1)
 moles PCl_3 needed = $4 \times 0.352 = 1.408/1.4 \text{ mol}$ (1)
 mass $\text{PCl}_3 = 1.4(08) \times 208.5 = 293.568 / 294 \text{ g} / 291.9 \text{ g}$ (1)

[7]

[13]

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