

*Honors Chemistry Midterm: Review Packet***PART 1**

1. Write in the number of significant figures for each of the following measurements:

3 3.00 mL      3 0.00290 g      4 50.00 m      2 0.070 kg      1 400 L

2. Using the following masses: 0.0562 kg, 124.213 g and 1635 mg.
- ← convert to grams*
- 
- The answer should be reported as:
- 
- g

*56.2g    124.213g    1.635g*

3. a) A 5.75 mL sample of mercury has a measured mass of 77.05 g. The density is
- 13.49/mL

$$D = \frac{m}{V} = \frac{77.05g}{5.75mL} =$$

- b) Mercury's accepted density is 13.53 g/mL. The % error in your measurement is:
- .9617%

$$\frac{|\text{actual} - \text{exp}|}{\text{actual}} \times 100 \quad \frac{|13.53g/mL - 13.49g/mL|}{13.53g/mL} \times 100$$

4. In the diagram below:



Is this person accurate? no - not in center

Is this person precise? yes - close together

5. Identify the following as a chemical change (CC), physical change (PC), chemical property (CP), or physical property (PP):

a) burning wood CC      b) mass of wood PP  
c) flammability of wood CP      d) chopping up wood PC

# Mcps Honors Chemistry B Review Packet Answers

**Yi-Tong Ma**



**Mcps Honors Chemistry B Review Packet Answers:**

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### **Table of Contents Mcps Honors Chemistry B Review Packet Answers**

1. Understanding the eBook Mcps Honors Chemistry B Review Packet Answers
  - The Rise of Digital Reading Mcps Honors Chemistry B Review Packet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Mcps Honors Chemistry B Review Packet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mcps Honors Chemistry B Review Packet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mcps Honors Chemistry B Review Packet Answers
  - Personalized Recommendations
  - Mcps Honors Chemistry B Review Packet Answers User Reviews and Ratings
  - Mcps Honors Chemistry B Review Packet Answers and Bestseller Lists

5. Accessing Mcps Honors Chemistry B Review Packet Answers Free and Paid eBooks
  - Mcps Honors Chemistry B Review Packet Answers Public Domain eBooks
  - Mcps Honors Chemistry B Review Packet Answers eBook Subscription Services
  - Mcps Honors Chemistry B Review Packet Answers Budget-Friendly Options
6. Navigating Mcps Honors Chemistry B Review Packet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Mcps Honors Chemistry B Review Packet Answers Compatibility with Devices
  - Mcps Honors Chemistry B Review Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mcps Honors Chemistry B Review Packet Answers
  - Highlighting and Note-Taking Mcps Honors Chemistry B Review Packet Answers
  - Interactive Elements Mcps Honors Chemistry B Review Packet Answers
8. Staying Engaged with Mcps Honors Chemistry B Review Packet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mcps Honors Chemistry B Review Packet Answers
9. Balancing eBooks and Physical Books Mcps Honors Chemistry B Review Packet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mcps Honors Chemistry B Review Packet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Mcps Honors Chemistry B Review Packet Answers
  - Setting Reading Goals Mcps Honors Chemistry B Review Packet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcps Honors Chemistry B Review Packet Answers
  - Fact-Checking eBook Content of Mcps Honors Chemistry B Review Packet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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